


A1 States of matter


1. a) i) Ammonium chloride (1)
 ii) Diffusion (1)
 Diffusion is the movement of particles from an area of higher concentration to an area of lower concentration until the particles are evenly distributed. (2)
 iii) The ammonia solution gave off ammonia gas and the hydrochloric acid gave off hydrogen chloride gas. The particles of the two gases diffused along the tube, collided and reacted to form ammonium chloride. (3)
- b) i)  (1)
 ii) The distilled water had a higher water content than the cytoplasm inside the paw paw cells and the cell membranes were differentially permeable. The water molecules were able to move through the membranes into the cells causing the cells to swell and the strips to increase in volume. (3)
 iii) Osmosis (1)
- c) Sodium chloride draws water out of the cells of the fish and any micro-organisms by osmosis. Water is, therefore, not available for chemical reactions in the cells, some of which cause decay and micro-organisms can't grow. (2)

2. a)

Property	Solid	Liquid	Gas
Volume		definite	variable; the volume is the same as the entire container
Arrangement of particles	packed closely together, usually in a regular way		randomly arranged with large spaces between
Energy of particles	have very small amounts of kinetic energy	have medium amounts of kinetic energy	

(6)

- b) i) The particles have large spaces between them, so they can be easily pushed closer together. (1)
 ii) The particles are packed tightly together with very little empty space between. (1)
 iii) The particles move around rapidly and have weak forces of attraction between them, so they spread out to fill any available space in the container. (2)

- c) i) A: Melting B: Boiling or evaporation (4)
 C: Freezing D: Condensation (4)
 ii) Iodine or carbon dioxide or ammonium chloride (1)
 or naphthalene (1)
 d) i) Liquid (1)
 ii) 54 °C (1)
 iii)  (1)

A2 Mixtures and separations

1. a) i)

	Pure substance	Mixture
Composition		variable
Properties	fixed and constant	variable; the components retain their individual properties

 (3)
- ii) An element is a pure substance that cannot be broken down into any simpler substances by any ordinary chemical or physical means. A compound is a pure substance that contains two or more different types of element that are bonded together chemically in fixed proportions and in such a way that their properties have changed. (2)
- b) i) The particles in a suspension are larger than those in a colloid. (1)
 ii) The particles in a suspension settle if left undisturbed, whereas the particles in a colloid never settle. (1)
 iii) Light does not pass through a suspension, whereas most colloids scatter light. (1)
- c) i) A solution is a homogeneous mixture consisting of two or more components, one of which is usually a liquid. (1)
 ii) Solubility is the mass of solute that will saturate 100 g of solvent at a given temperature. (1)
- d) i) 17 g per 100 g water (1)
 ii) 45 °C (1)
 iii) At 76 °C, 54 g of Z saturate 100 g water
 At 36 °C, 12.5 g of Z saturate 100 g water
 \therefore mass of Z crystallising out of a saturated solution containing 100 g water = $54 - 12.5$ g = 41.5 g (3)
- iv) At 62 °C, 39 g of Z saturate 100 g water
 \therefore at 62 °C, 45 g of Z saturate $\frac{100}{39} \times 45$ g water = 115.4 g of water (2)
- v) At 55 °C, 33 g of Z saturate 100 g water
 \therefore at 55 °C, $\frac{33}{100} \times 350$ g of Z saturate 350 g water = 115.5 g of Z (2)

States Of Matter Chemistry Workbook Answers

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Brutlag, Michelle Rose Gilman, Brian
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Table of Contents States Of Matter Chemistry Workbook Answers

1. Understanding the eBook States Of Matter Chemistry Workbook Answers
 - The Rise of Digital Reading States Of Matter Chemistry Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying States Of Matter Chemistry Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an States Of Matter Chemistry Workbook Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from States Of Matter Chemistry Workbook Answers
 - Personalized Recommendations
 - States Of Matter Chemistry Workbook Answers User Reviews and Ratings
 - States Of Matter Chemistry Workbook Answers and Bestseller Lists

5. Accessing States Of Matter Chemistry Workbook Answers Free and Paid eBooks
 - States Of Matter Chemistry Workbook Answers Public Domain eBooks
 - States Of Matter Chemistry Workbook Answers eBook Subscription Services
 - States Of Matter Chemistry Workbook Answers Budget-Friendly Options
6. Navigating States Of Matter Chemistry Workbook Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - States Of Matter Chemistry Workbook Answers Compatibility with Devices
 - States Of Matter Chemistry Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of States Of Matter Chemistry Workbook Answers
 - Highlighting and Note-Taking States Of Matter Chemistry Workbook Answers
 - Interactive Elements States Of Matter Chemistry Workbook Answers
8. Staying Engaged with States Of Matter Chemistry Workbook Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers States Of Matter Chemistry Workbook Answers
9. Balancing eBooks and Physical Books States Of Matter Chemistry Workbook Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection States Of Matter Chemistry Workbook Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine States Of Matter Chemistry Workbook Answers
 - Setting Reading Goals States Of Matter Chemistry Workbook Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of States Of Matter Chemistry Workbook Answers
 - Fact-Checking eBook Content of States Of Matter Chemistry Workbook Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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