

## Section 8.2

### The Nature of Covalent Bonding

- OBJECTIVES:

- Describe how oxygen atoms are bonded in ozone.

# The Nature Of Covalent Bonding Lesson 8 2

**RM Cervero**



## **The Nature Of Covalent Bonding Lesson 8 2:**

**The Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Pauling, 1960 Thorough discussion of the various types of bonds their relative natures and the structure of molecules and crystals **Symmetry and Physical Properties of Crystals** Cécile Malgrange, Christian Ricolleau, Michel Schlenker, 2014-12-04 Crystals are everywhere from natural crystals minerals through the semiconductors and magnetic materials in electronic devices and computers or piezoelectric resonators at the heart of our quartz watches to electro optical devices Understanding them in depth is essential both for pure research and for their applications This book provides a clear thorough presentation of their symmetry both at the microscopic space group level and the macroscopic point group level The implications of the symmetry of crystals for their physical properties are then presented together with their mathematical description in terms of tensors The conditions on the symmetry of a crystal for a given property to exist then become clear as does the symmetry of the property The geometrical representation of tensor quantities or properties is presented and its use in determining important relationships emphasized An original feature of this book is that most chapters include exercises with complete solutions This allows readers to test and improve their understanding of the material The intended readership includes undergraduate and graduate students in materials science and materials related aspects of electrical and optical engineering researchers involved in the investigation of the physical properties of crystals and the design of applications based on crystal properties such as piezoelectricity electro optics optical activity and all those involved in the characterization of the structural properties of materials Taylor & Francis Group, 2010-12-31 [CBSE Class 10 Term 2 Chapterwise Question Bank Science by Career Point, Kota](#) Career Point Kota, 2022-01-18 Strictly as per the Term II syllabus for Board 2022 Exams March April Includes Questions of the both Objective Subjective Types Questions Objective Questions based on new typologies introduced by the board Stand Alone MCQs MCQs based on Assertion Reason Case based MCQs Subjective Questions includes Short Long Answer Types Questions Include Questions from CBSE official Question Bank released in April 2021 Chapter wise Tests 2 Full Syllabus Practice Papers *Materials Design Inspired by Nature* Peter Fratzl, John Dunlop, Richard Weinkamer, 2013-05-24 The inner architecture of a material can have an astonishing effect on its overall properties and is vital to understand when designing new materials Nature is a master at designing hierarchical structures and so researchers are looking at biological examples for inspiration specifically to understand how nature arranges the inner architectures for a particular function in order to apply these design principles into man made materials Materials Design Inspired by Nature is the first book to address the relationship between the inner architecture of natural materials and their physical properties for materials design The book explores examples from plants the marine world arthropods and bacteria where the inner architecture is exploited to obtain specific mechanical optical or magnetic properties along with how these design principles are used in man made products Details of the experimental methods used to investigate hierarchical structures are also given

Written by leading experts in bio inspired materials research this is essential reading for anyone developing new materials

**Chemistry of Polysaccharides** Gennady Zaikov, 2005-04-18 This monograph summarizes scientific achievements in the field of polysaccharide chemistry performed in the last decade For the first time the particular features of polysaccharide structure physical chemical properties and biochemical transformations in their interrelations are considered as well as the questions of polysaccharide modification along the whole hierarchical chain including starch mono di oligo and polysaccharides chitin chitozan and cellulose Polysaccharide interactions with diverse solvents and low molecular macrocyclic and high molecular substances in solutions are scrutinized Several aspects of mechanical chemical modification of polysaccharides in solid state and in gel are observed Modern concepts of enzyme degradation of polysaccharides and technologies of natural polymers processing are analyzed The book is intended for wide range of readers students post graduates engineers and scientists engaged in polymer chemistry organic and physical chemistry Seven Clues to the

Origin of Life Alexander Graham Cairns-Smith, 1990-09-13 The mysteries surrounding the origins of life on earth are written in detective story fashion by a world famous scientist in this popular version of Genetic Takeover originally published in 1982

**Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **An Introduction to Mineral Sciences** A. Putnis, 1992-10-22 The subject of mineralogy is moving away from the traditional systematic treatment of mineral groups toward the study of the behaviour of minerals in relation to geological processes A knowledge of how minerals respond to a changing geological environment is fundamental to our understanding of many dynamic earth processes By adopting a materials science approach An Introduction to Mineral Sciences explains the principles underlying the modern study of minerals discussing the behaviour of crystalline materials with changes in temperature pressure and chemical environment The concepts required to understand mineral behaviour are often complex but are presented here in simple non mathematical terms for undergraduate mineralogy students After introductory chapters describing the principles of diffraction imaging and the spectroscopic methods used to study minerals the structure and behaviour of the main groups of rock forming minerals are covered and the role of defects in the deformation and transformation of a mineral are explained The energy changes and the rate of transformation processes are introduced using a descriptive approach rather than attempting a complete and rigorous treatment of the thermodynamics and kinetics Examples and case histories from a range of mineral groups are set in an earth science context such that the emphasis of this book is to allow the student to develop an intuitive understanding of the structural principles controlling the behaviour of minerals **Trace Metals in Aquatic Systems** Robert P. Mason, 2013-02-20 This book provides a detailed examination of the concentration form and cycling of trace metals and metalloids through the aquatic biosphere and has sections dealing with the atmosphere the ocean lakes and rivers It discusses exchanges at the water interface air water and sediment water and the major drivers of the cycling concentration and form of trace metals in aquatic

systems The initial chapters focus on the fundamental principles and modelling approaches needed to understand metal concentration speciation and fate in the aquatic environment while the later chapters focus on specific environments with case studies and research highlights Specific examples deal with metals that are of particular scientific interest such as mercury iron arsenic and zinc and the book deals with both pollutant and required nutrient metals and metalloids The underlying chemical principles controlling toxicity and bioavailability of these elements to microorganisms and to the aquatic food chain are also discussed Readership Graduate students studying environmental chemistry and related topics as well as scientists and managers interested in the cycling of trace substances in aqueous systems Additional resources for this book can be found at [www.wiley.com/go/mason/tracemetals](http://www.wiley.com/go/mason/tracemetals) Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance **Advanced Physiology and Pathophysiology** Nancy Tkacs, Nancy C. Tkacs, Linda Herrmann, 2024-03-22 2020 AJN Book of the Year Award Winner Advanced Practice Nursing Concise and well organized this advanced physiology and pathophysiology text promotes deep understanding of key pathophysiological concepts and relates them to major disorders commonly seen in practice It is distinguished by its stellar organization and writing that clarifies difficult to understand disease mechanisms The second edition offers several new features that add to its value in preparing the student clinician to care for patients with a broad variety of disorders effectively The second edition retains the inclusive language and conceptual organization central to the appeal and usefulness of the first edition Coverage of new scientific advances update the original disease descriptions With the aim of preparing students to be well informed and confident decision makers in primary care settings the text provides the key knowledge required to master pharmacology for prescriptive practice Addressing both healthy organ function and disease associated changes it details and illustrates the cellular structure and function of each organ system and mechanisms of associated major clinical disorders It examines the reasons patients often present with particular symptoms the rationale for ordering specific diagnostic tests and interpretation of the results and common management strategies that proceed from the underlying pathology The text is replete with case presentations to illustrate concepts over 500 images key points at the end of each chapter to reinforce knowledge and a glossary defining correct terminology Comprehensive instructor resources

accompany the text New to the Second Edition Standardized template for pediatric and gerontology content for clarity Expanded coverage of neonatal and pediatric development and vulnerabilities Methylation patterns of cell free DNA epigenetics in oncology practice Expanded coverage of fluid and electrolyte balance Physiology of pregnancy labor and delivery Increased content on the social determinants of health Introduction of the exposome concept in human disease New content on biological and psychosocial aspects of human sexual development and variations relevant to LGBTQ centered care COVID 19 related concerns Key Features Includes comprehensive lifespan considerations with key insights from specialists in pediatric and geriatric pathophysiology Integrates critical thinking questions and case studies to promote discussion and information synthesis Provides unique Bridge to Clinical Practice in each chapter to translate science to practical patient care Includes more than 500 images to illustrate complex scientific concepts Summarizes content with key points at the end of each chapter **Principles and Measurements in Environmental Biology** F I Woodward,J E Sheehy,2017-05-04

**Principles and Measurements in Environmental Biology** aims to provide an understanding of some important physical principles and their application in biology The book also aims to describe how instruments utilizing these principles can be used to measure biological and environmental processes and their interactions This book covers the effects of the environment on biological organisms the application of theories of radiation kinetic theory gas laws and diffusion in biology and water and its properties The relation of plants with atmosphere near the ground is also discussed This book also presents sampling techniques the computation of errors used in the interpretation of data the use of different devices and data gathering and its practical applications This text is for students researchers and professionals and experts in biology who wish to understand the mentioned principles in physics its mathematical aspects and their applications in the field

**Chemical Fundamentals of Geology and Environmental Geoscience** Robin Gill,2015-01-27 Chemical principles are fundamental to the Earth sciences and geoscience students increasingly require a firm grasp of basic chemistry to succeed in their studies The enlarged third edition of this highly regarded textbook introduces the student to such geo relevant chemistry presented in the same lucid and accessible style as earlier editions but the new edition has been strengthened in its coverage of environmental geoscience and incorporates a new chapter introducing isotope geochemistry The book comprises three broad sections The first Chapters 1 4 deals with the basic physical chemistry of geological processes The second Chapters 5 8 introduces the wave mechanical view of the atom and explains the various types of chemical bonding that give Earth materials their diverse and distinctive properties The final chapters 9 11 survey the geologically relevant elements and isotopes and explain their formation and their abundances in the cosmos and the Earth The book concludes with an extensive glossary of terms appendices cover basic maths explain basic solution chemistry and list the chemical elements and the symbols units and constants used in the book **Schaum's Outline of Beginning Chemistry (EBOOK)**

David E. Goldberg,2013-09-27 Tough Test Questions Missed Lectures Not Enough Time Fortunately there s Schaum s This all

in one package includes more than 650 fully solved problems examples and practice exercises to sharpen your problem solving skills Plus you will have access to 16 detailed videos featuring Chemistry instructors who explain the most commonly tested concepts it's just like having your own virtual tutor You'll find everything you need to build confidence skills and knowledge for the highest score possible More than 40 million students have trusted Schaum's to help them succeed in the classroom and on exams Schaum's is the key to faster learning and higher grades in every subject Each Outline presents all the essential course information in an easy to follow topic by topic format You also get hundreds of examples solved problems and practice exercises to test your skills This Schaum's Outline gives you 673 fully solved problems Hundreds of examples with explanations of chemistry concepts Support for all the major textbooks for beginning chemistry courses Fully compatible with your classroom text Schaum's highlights all the important facts you need to know Use Schaum's to shorten your study time and get your best test scores Schaum's Outlines Problem Solved

**Chemistry of Aluminium, Gallium, Indium and Thallium** A.J. Downs, 1993-05-31 Boron has all the best tunes That may well be the first impression of the Group 13 elements The chemical literature fosters the impression not only in the primary journals but also in a steady outflow of books focussing more or less closely on boron and its compounds The same preoccupation with boron is apparent in the coverage received by the Group 13 elements in the comprehensive and regularly updated volume of the Gmelin Handbook Yet such an imbalance cannot be explained by any inherent lack of variety interest or consequence in the heavier elements Aluminium is the most abundant metal in the earth's crust in the industrialised world the metal is second only to iron in its usage and its compounds can justifiably be said to touch our lives daily to the potential detriment of those and other lives some would argue From being chemical curios gallium and indium have now gained considerably prominence as sources of compound semiconductors like gallium arsenide and indium antimonide Nor is there any want of incident in the chemistries of the heavier Group 13 elements In their redox coordination and structural properties there is to be found music indeed notable not always for its harmony but invariably for its richness and variety This book seeks to redress the balance with a definitive wide ranging and up to date review of the chemistry of the Group 13 metals aluminium gallium indium and thallium

**Theoretical and Quantum Chemistry at the Dawn of the 21st Century** Tanmoy Chakraborty, Ramon Carbo-Dorca, 2018-06-19 This volume edited by a well known specialist in the field of theoretical chemistry gathers together a selection of papers on theoretical chemistry within the themes of mathematical computational and quantum chemistry The authors present a rich assembly of some of the most important current research in the field of quantum chemistry in modern times In Quantum Chemistry at the Dawn of the 21st Century the editors aim to replicate the tradition of the fruitful Girona Workshops and Seminars held at the University of Girona Italy annually for many years which offered important scientific gatherings focusing on quantum chemistry This volume like the workshops showcases a large variety of quantum chemical contributions from different points of view from some of the leading scientists in the field today This unique volume does not pretend to provide a complete

overview of quantum chemistry but it does provide a broad set of contributions by some of the leading scientists on the field under the expert editorship of two leaders in the field

**X-Ray Charge Densities and Chemical Bonding** Philip Coppens, 1997-05-08 This book deals with the electron density distribution in molecules and solids as obtained experimentally by X ray diffraction It is a comprehensive treatment of the methods involved and the interpretation of the experimental results in terms of chemical bonding and intermolecular interactions Inorganic and organic solids as well as metals are covered in the chapters dealing with specific systems As a whole this monograph is especially appealing because of its broad interface with numerous disciplines Accurate X ray diffraction intensities contain fundamental information on the charge distribution in crystals which can be compared directly with theoretical results and used to derive other physical properties such as electrostatic moments the electrostatic potential and lattice energies which are accessible by spectroscopic and thermodynamic measurements Consequently the work will be of great interest to a broad range of crystallographers and physical scientists

**Physical Pharmacy** Dr. U. B. Hadkar, 2007-07

*Linus Pauling - Selected Scientific Papers (In 2 Volumes) - Volume 1* Barclay Kamb, Linda Pauling Kamb, Carl Alexander Kamb, Peter Jeffress Pauling, Linus Pauling Jr, 2001-11-02 Linus Pauling wrote a stellar series of over 800 scientific papers spanning an amazing range of fields some of which he himself initiated This book is a selection of the most important of his writings in the fields of quantum mechanics chemical bonding covalent ionic metallic and hydrogen bonding molecular rotation and entropy protein structure hemoglobin molecular disease molecular evolution the antibody mechanism the molecular basis of anesthesia orthomolecular medicine radiation chemistry biology and nuclear structure Through these papers the reader gets a fresh unfiltered view of the genius of Pauling s many contributions to chemistry chemical physics molecular biology and molecular medicine



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## The Nature Of Covalent Bonding Lesson 8 2 Introduction

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## **The Nature Of Covalent Bonding Lesson 8 2 :**

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web oxidation reduction redox reactions oxidation  $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^{-}$  reduction  $\text{MnO}_4^{-} + 4\text{H}^{+} + 3\text{e}^{-} \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$  l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients  $\text{Cu} : \text{MnO}_4^{-} : 4$  is

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2^{-} + \text{Al} + \text{OH}^{-} \rightarrow \text{AlO}_2^{-} + \text{NH}_3$  practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^{-}$  previous topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^{-} + \text{SO}_2 + 2\text{H}_2\text{O} \rightarrow \text{Mn}^{2+} + \text{HSO}_4^{-}$  what will be the stoichiometric coefficient of  $\text{HSO}_4^{-}$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web may 25 2021 balancing redox equations using half reactions another way to balance redox reactions is by the half reaction method this technique involves breaking an equation into its two separate components the oxidation reaction and

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} \rightarrow 2\text{HCN} + 2\text{OH}^{-}$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-} \rightarrow 4\text{HCN} + 2\text{OH}^{-}$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web apr 17 2013 in order to balance redox reactions well you must first be able to assign oxidation numbers well oxidation the loss of electrons resulting in a more positively charged species reduction the gain of electrons resulting in a

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web a redox equation can be balanced using the following stepwise procedure 1 divide the equation into two half reactions 2 balance each half reaction for mass and charge 3 equalize the number of electrons transferred in each half reaction 4

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic solution created by jay

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web bonus problem here is a balanced net ionic equation 3br<sub>2</sub> + 6oh<sup>-</sup> → 5br<sup>-</sup> + bro<sup>-</sup> + 3h<sub>2</sub>o the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web mar 13 2023 balancing redox equations by the ion electron method there are two principal methods for balancing redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s → ni a + 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}^{2+}$   $\text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}^{2+}$   $2\text{Cr}^{3+}$   $7\text{H}_2\text{O}$   $\text{Fe}^{2+}$   $\text{Fe}^{3+}$   $\text{e}^-$  2 equalize the electrons  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}^{2+}$   $2\text{Cr}^{3+}$   $7\text{H}_2\text{O}$   $6\text{Fe}^{2+}$   $6\text{Fe}^{3+}$   $6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}^{2+}$   $\text{NO}_3^-$   $\text{SO}_4^{2-}$   $\text{MnO}_2$   $\text{Cr}^{3+}$   $\text{NO}_3^-$   $\text{SO}_2$   $\text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2$   $\text{Cr}^{2+}$   $\text{TeO}_3$   $\text{N}_2\text{O}_4$   $\text{O}_2$   $\text{Cr}^{3+}$   $\text{Te}$  10  $\text{ReO}_4^-$  11  $\text{PbO}_2$   $\text{IO}_3^-$   $\text{I}_2$   $\text{Pb}^{2+}$  12  $\text{AsO}_3^{3-}$   $\text{IO}_3^-$   $\text{H}_2\text{AsO}_4^-$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}^+$   $\text{O}^{2-}$  and  $\text{e}^-$  in this order balance the atoms in the equation apart from  $\text{O}$  and  $\text{H}$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$   $2\text{H}_2\text{NO}_2$   $\text{C}_6\text{H}_6\text{O}_6$   $2\text{H}_2\text{O}_2$   $\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  2 and  $\text{C}_6\text{H}_8\text{O}_6$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and  $\text{NO}_2$  c h and

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