Basic Chemistry Study Guide 2

Hydrogen bonds

- Form between two polar molecules covalent bonds
- A hydrogen's &+ region is attracted to a &- region in another molecule.
- Form between water molecules (Where are the &+ and &- regions in a water molecule?) hydrogen bonds, at the top of the symbol
- Each hydrogen bond is very weak (can be easily formed and broken).
- There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
- One of these properties is the ability to be a solvent for other polar molecules.
- Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

- Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
- Water is located inside and outside of our cells.
- 3. Body fluids are solutions
- Define the terms solution, solvent, and solute. Solvent a substance capable of dissolving other polar molecules and ionic bonds, solutes, molecules dissolved in the solvent, solution homogeneous moture.

3. How substances dissolve in water

- When a solute dissolves, it becomes surrounded by water molecules.
- 2. For a solute to dissolve, it must attract water molecules.
- What kinds of substances attract water molecules? Sugars, nucleic acids, and amino acids
- When a solute dissolves it forms hydrogen bonds with water molecules.
- Define the terms ion, cation, and anion. Ion atom with a charge, cation; positively charged ion, anions; negatively charged ion.
- Define the terms hydrophilic and hydrophobic Hydrophilic attracted to water, hydrophobic, a kaid of water
- What types of molecules are hydrophilic and what types are hydrophobic? Nonpolar hydrophobic, polar hydrophilic

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the manufacturer's default injection timing is 6 5 deg before tdc the only major changes from the 1kd ftv to the 2kd ftv were the bore and the stroke in 2005 denso corporation introduced the 3rd generation of common rail system

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