

# WRITING CHEMICAL EQUATIONS

| Name _____                                                                                                                                                                                                   | Section _____ | Date _____ |
|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------------|------------|
| <b>Writing Equations</b>                                                                                                                                                                                     |               |            |
| <b>Directions:</b> Write and balance each chemical equation. You will need to write the products. Include phase labels with each reactant and product. You may need to refer to a table of solubility rules. |               |            |
| 1) Aqueous sulfuric acid reacts with aqueous potassium hydroxide                                                                                                                                             |               |            |
| 2) Aqueous potassium carbonate reacts with aqueous hydrochloric acid                                                                                                                                         |               |            |
| 3) Hydrogen and oxygen gases combine to form water vapor                                                                                                                                                     |               |            |

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| 2) Aqueous potassium carbonate reacts with aqueous hydrochloric acid                                                                                                                                         |               |            |
| 3) Hydrogen and oxygen gases combine to form water vapor                                                                                                                                                     |               |            |
| 4) Methane gas (carbon tetrahydride) reacts with chlorine gas to yield chloroform (CH <sub>3</sub> Cl) gas and hydrochloric acid gas                                                                         |               |            |
| 5) Aqueous ammonium sulfide reacts with aqueous magnesium sulfate to yield solid magnesium sulfide and aqueous ammonium sulfate                                                                              |               |            |
| 6) Mercury (II) nitrate reacts with aqueous iron(III) iodide to form aqueous iron (III) nitrate and solid mercury (II) iodide                                                                                |               |            |
| 7) Aqueous strontium hydroxide reacts with aqueous iron (III) nitrate to produce solid iron hydroxide and aqueous strontium nitrate                                                                          |               |            |
| 8) Solid zinc carbonate reacts with aqueous potassium chloride. An aqueous solution of zinc chloride and potassium carbonate is formed.                                                                      |               |            |
| 9) Aqueous ammonium chromate reacts with aqueous barium sulfide react to form aqueous ammonium sulfide and solid barium chromate.                                                                            |               |            |
| 10) Aqueous lithium oxalate reacts with aqueous barium sulfate to form aqueous lithium sulfate and solid barium oxalate                                                                                      |               |            |

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A CHEMISTRY WORKSHEET

10 PROBLEMS

WORKED ANSWER KEY INCLUDED

# Topic 3 Formulas And Equations Answer Key

**SA Dillow**



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