CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY

 Which of the following are weak electrolytes? HCI, HC.H.O., NH., KCI A) HCL. KCI D) HC,H,O,, NH, B) HCl, HC, H, O, , NH, , KCl E) HCI, HC, H, O, , KCI C) HC₂H₃O₂, KCI Answer: D Dim 2

2) What are the spectator ions in the reaction between KOH (aq) and HNO, (aq)?

A) K and H C) K'and NOs" E) OHI only. B) H'and OH" D) H'and NO. Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of CaCO, when aqueous solutions of

Na . CO, and CaCl are mixed is .

A) 2Na*(aq)+CO,²-(aq) → Na, CO, (aq) ET N

Na ,CO ,(aq) + CaCl , (aq) → 2NaCl (aq) + CaCO ,(s) B) 2 Na*(aq) + 2Cl*(aq) → 2NaCl (aq)

C) Na*(aq)+Cl*(aq) → NaCl (aq)

D) $Ca^{2+}(aq) + CO_1^{2-}(aq) \longrightarrow CaCO_1(s)$

American: D DMT = 3

4) Which of the following are strong acids? HI, HNO., HF, HBr

A) HF, HBr C) HI, HF, HBr

B) HI, HNO, HF, HBr D) HNO., HF, HBr

Australia E Distr. 2

5) A neutralization reaction between an acid and a metal hydroxide produces

A) water and a salt C) oxygen gas E) amormomia.

B) hydrogen gas D) sodium hydroxide

Amswers A. Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na B) Au C) Fe D) Ca Ammuer: B Dan: 3

7) Based on the equations below, which metal is the most active?

 $Pb(NO_+)_+(aq) + Ni(s) \rightarrow Ni(NO_+)_+(aq) + Pb(s)$

 $Pb(NO_{+})_{+}(aq) + Ag(s) \rightarrow No reaction$

 $Cu(NO_+)_*(aq) + Ni(s) \rightarrow Ni(NO_+)_*(aq) + Cu(s)$

A) Ni C) Cu B) Ast Answer: A Diff: 3

D) Pb

E) N

E) Ag

E) HI, HNO, HBr

 What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL. of 0.100 M KCI?

A) 0.100 B) 0.0500 Alternatives Al. Diff: 3

C) 0.0333

D) 0.0250

E) 125

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