

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

Speed of light = (frequency)(wavelength)

$$c = f\lambda$$

Energy = (Planck's constant)(frequency)

$$E_{\text{photon}} = hf$$

Energy = $\frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})}$

$$E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

Total pressure of a gas = $\left(\frac{\text{sum of the partial pressures}}{\text{of the component gases}} \right)$

$$P_T = P_1 + P_2 + P_3 + \dots$$

(Pressure)(volume) = (moles)(ideal gas constant)(temperature)

$$PV = nRT$$

$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})}$

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

(Initial pressure)(initial volume) = (final pressure)(final volume)

$$P_1V_1 = P_2V_2$$

$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})}$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})}$

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

Molarity = $\frac{\text{moles of solute}}{\text{liter of solution}}$

$$M = \frac{\text{mol}}{\text{L}}$$

Ionization constant of water = $\left(\frac{\text{hydrogen ion}}{\text{concentration}} \right) \left(\frac{\text{hydroxide ion}}{\text{concentration}} \right)$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$\left(\frac{\text{Volume of}}{\text{solution 1}} \right) \left(\frac{\text{molarity of}}{\text{solution 1}} \right) = \left(\frac{\text{volume of}}{\text{solution 2}} \right) \left(\frac{\text{molarity of}}{\text{solution 2}} \right)$

$$V_1M_1 = V_2M_2$$

pH = -logarithm (hydrogen ion concentration)

$$\text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

Heat gained or lost = (mass) $\left(\frac{\text{specific}}{\text{heat}} \right) \left(\frac{\text{change in}}{\text{temperature}} \right)$

$$Q = mc_p\Delta T$$

Enthalpy of reaction = $\left(\frac{\text{enthalpy}}{\text{of products}} \right) - \left(\frac{\text{enthalpy}}{\text{of reactants}} \right)$

$$\Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

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E Durkheim



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