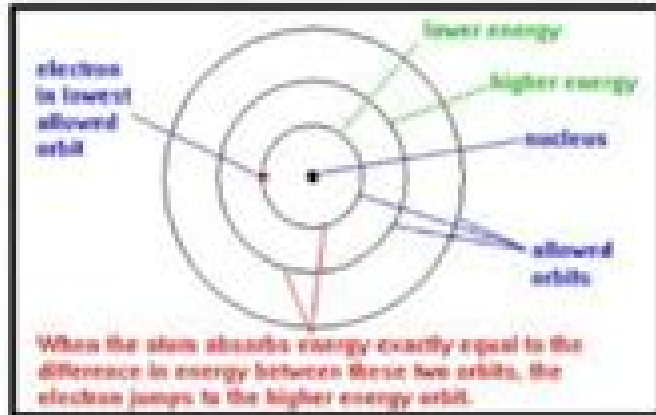


The Bohr Model

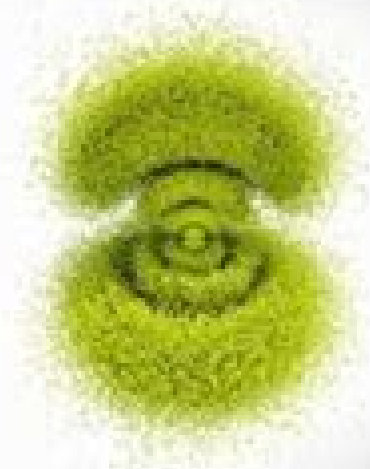
A one-dimensional model, one quantum number (n) to describe the distribution of electrons in the atom.



3



what
atoms
really
look like?



4



Quantum electron configuration

Angular (l): describes the shape of the orbital.

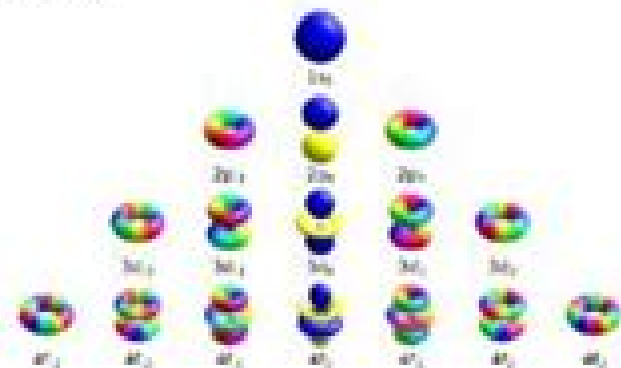
- Orbitals have shapes that are best described as spherical ($l = 0$), polar ($l = 1$), or cloverleaf ($l = 2$).
- They can even take on more complex shapes as the value of the angular quantum number becomes larger.

Angular Quantum Number, l	Name of Orbital	Shape
0	s	Spherical
1	p	Dumbbell
2	d	Complex (clover or dumbbell)
3	f	More complex/multiple lobes

Quantum electron configuration

Magnetic (m_l): describes the orientation in space

- There is only one way in which a sphere ($l = 0$) can be oriented in space.
- l is called the magnetic quantum number because the effect of different orientations of orbitals was first observed in the presence of a magnetic field.



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