

## Chemical Equilibrium

### Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when
- the rate of the forward reaction equals that of the backward reaction.
  - the concentration of reactants and products are equal
  - the activation energies of the forward and backward reactions are equal
  - the reaction involves no enthalpy change.
2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

- high temperature, low pressure
  - high temperature, high pressure
  - low temperature, high pressure
  - low temperature, low pressure.
3.  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
- In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because
- the ammonia is constantly removed
  - an iron catalyst is used
  - high pressure is used
  - the temperature is held at above  $500^\circ\text{C}$ .

4. The equation refers to the preparation of methanol from synthesis gas.
- $$\text{CO}(\text{g}) + 2\text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$$
- $\Delta H = -91 \text{ kJ mol}^{-1}$

The formation of methanol is favoured by

- high pressure and low temperature
  - high pressure and high temperature
  - low pressure and low temperature
  - low pressure and high temperature.
5.  $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$
- $\Delta H = -46 \text{ kJ mol}^{-1}$

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

	Temperature	Pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

6. When a reversible chemical reaction is at equilibrium,
- the concentration of reactants and products remain equal
  - the forward reaction is unable to continue
  - the concentration of reactants and products remain constant
  - the forward and reverse reactions proceed at different rates.

# Static Equilibrium Multiple Choice Questions

**S Marginson**



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