

Chemical Equilibrium

Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when
- A. the rate of the forward reaction equals that of the backward reaction.
 - B. the concentration of reactants and products are equal
 - C. the activation energies of the forward and backward reactions are equal
 - D. the reaction involves no enthalpy change.
2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

- A. high temperature, low pressure
 - B. high temperature, high pressure
 - C. low temperature, high pressure
 - D. low temperature, low pressure.
3. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
- In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because
- A. the ammonia is constantly removed
 - B. an iron catalyst is used
 - C. high pressure is used
 - D. the temperature is held at above 500°C .

4. The equation refers to the preparation of methanol from synthesis gas.
- $$\text{CO}(\text{g}) + 2\text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$$
- $$\Delta H = -91 \text{ kJ mol}^{-1}$$

The formation of methanol is favoured by

- A. high pressure and low temperature
 - B. high pressure and high temperature
 - C. low pressure and low temperature
 - D. low pressure and high temperature.
5. $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$
- $$\Delta H = -46 \text{ kJ mol}^{-1}$$

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

	Temperature	Pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

6. When a reversible chemical reaction is at equilibrium,
- A. the concentration of reactants and products remain equal
 - B. the forward reaction is unable to continue
 - C. the concentration of reactants and products remain constant
 - D. the forward and reverse reactions proceed at different rates.

Static Equilibrium Multiple Choice Questions

S Marginson



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Table of Contents Static Equilibrium Multiple Choice Questions

1. Understanding the eBook Static Equilibrium Multiple Choice Questions
 - The Rise of Digital Reading Static Equilibrium Multiple Choice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Static Equilibrium Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Static Equilibrium Multiple Choice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Static Equilibrium Multiple Choice Questions
 - Personalized Recommendations
 - Static Equilibrium Multiple Choice Questions User Reviews and Ratings
 - Static Equilibrium Multiple Choice Questions and Bestseller Lists
5. Accessing Static Equilibrium Multiple Choice Questions Free and Paid eBooks
 - Static Equilibrium Multiple Choice Questions Public Domain eBooks
 - Static Equilibrium Multiple Choice Questions eBook Subscription Services
 - Static Equilibrium Multiple Choice Questions Budget-Friendly Options
6. Navigating Static Equilibrium Multiple Choice Questions eBook Formats

- ePub, PDF, MOBI, and More
- Static Equilibrium Multiple Choice Questions Compatibility with Devices
- Static Equilibrium Multiple Choice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Static Equilibrium Multiple Choice Questions
 - Highlighting and Note-Taking Static Equilibrium Multiple Choice Questions
 - Interactive Elements Static Equilibrium Multiple Choice Questions
- 8. Staying Engaged with Static Equilibrium Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Static Equilibrium Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Static Equilibrium Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Static Equilibrium Multiple Choice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Static Equilibrium Multiple Choice Questions
 - Setting Reading Goals Static Equilibrium Multiple Choice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Static Equilibrium Multiple Choice Questions
 - Fact-Checking eBook Content of Static Equilibrium Multiple Choice Questions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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