

Name: _____
Period: _____

The Atom

Ch. 18:1

Everything is made of **atoms**. Atoms are the smallest part of matter.
Atoms are made up of 3 subatomic particles (particles smaller than the atom): *electrons, protons, and neutrons*.

On the following diagram of an atom define the parts of the atom.

Negatively charged particles; outside the nucleus; can be gained or lost to and from other atoms; very small mass: 1/2000th of a proton

electrons

protons

Positively charged particles; in the nucleus; determine what element an atom is.

This model of the atom looks a lot like a solar system. The nucleus, which contain the protons and neutrons, in the center would be the sun. The **electrons** are the planets spinning around the nucleus.

Neutral particles in the nucleus; give mass to the atom, but not charge.

neutrons

nucleus

Center of the atom; contains protons and neutrons.

Count the protons to tell what element this is:

of Protons: _____ Element: _____

John Dalton in 1808 published a theory of the atom that had these important points:

- All atoms of a particular element are the same.
- Atoms of different elements have different properties, mass, and chemical reactivity.
- Atoms are not changed by chemical reactions, just rearranged in order or number.

Atoms, Molecules, and Compounds

Atoms combine into **molecules**.
O is an atom; O₂ is a molecule; both are oxygen.

Molecules are made up of two or more atoms.

If two different atoms combine they make **compounds**.
H₂O is a compound; O₂ is a molecule.

Compounds are made up of two or more elements.

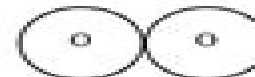
Water molecule—
a compound (H₂O)

2 hydrogens



1 oxygen

Oxygen molecule—
an element (O₂)



2 oxygens

Atom, molecule or compound?

NaCl — _____

Cl₂ — _____

Na — _____

What elements are these?

Na — _____

Cl — _____

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Table of Contents Stephen Murray Ch 13 Key Sounds

1. Understanding the eBook Stephen Murray Ch 13 Key Sounds
 - The Rise of Digital Reading Stephen Murray Ch 13 Key Sounds
 - Advantages of eBooks Over Traditional Books
2. Identifying Stephen Murray Ch 13 Key Sounds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stephen Murray Ch 13 Key Sounds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stephen Murray Ch 13 Key Sounds

- Personalized Recommendations
 - Stephen Murray Ch 13 Key Sounds User Reviews and Ratings
 - Stephen Murray Ch 13 Key Sounds and Bestseller Lists
5. Accessing Stephen Murray Ch 13 Key Sounds Free and Paid eBooks
 - Stephen Murray Ch 13 Key Sounds Public Domain eBooks
 - Stephen Murray Ch 13 Key Sounds eBook Subscription Services
 - Stephen Murray Ch 13 Key Sounds Budget-Friendly Options
 6. Navigating Stephen Murray Ch 13 Key Sounds eBook Formats
 - ePub, PDF, MOBI, and More
 - Stephen Murray Ch 13 Key Sounds Compatibility with Devices
 - Stephen Murray Ch 13 Key Sounds Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stephen Murray Ch 13 Key Sounds
 - Highlighting and Note-Taking Stephen Murray Ch 13 Key Sounds
 - Interactive Elements Stephen Murray Ch 13 Key Sounds
 8. Staying Engaged with Stephen Murray Ch 13 Key Sounds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stephen Murray Ch 13 Key Sounds
 9. Balancing eBooks and Physical Books Stephen Murray Ch 13 Key Sounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stephen Murray Ch 13 Key Sounds
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stephen Murray Ch 13 Key Sounds
 - Setting Reading Goals Stephen Murray Ch 13 Key Sounds
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stephen Murray Ch 13 Key Sounds

- Fact-Checking eBook Content of Stephen Murray Ch 13 Key Sounds
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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