

Chemistry Study Guide

Periodic Table of the Elements

<div>1 H</div>	<div>2 He</div>																	<div>18 Ar</div>	<div>19 K</div>	<div>20 Ca</div>											<div>36 Kr</div>	<div>37 Rb</div>	<div>38 Sr</div>											<div>54 Xe</div>	<div>55 Cs</div>	<div>56 Ba</div>															<div>86 Rn</div>																															
<div>3 Li</div>	<div>4 Be</div>																	<div>12 Mg</div>	<div>13 Al</div>	<div>14 Si</div>	<div>15 P</div>	<div>16 S</div>	<div>17 Cl</div>	<div>18 Ar</div>	<div>19 K</div>	<div>20 Ca</div>	<div>21 Sc</div>	<div>22 Ti</div>	<div>23 V</div>	<div>24 Cr</div>	<div>25 Mn</div>	<div>26 Fe</div>	<div>27 Co</div>	<div>28 Ni</div>	<div>29 Cu</div>	<div>30 Zn</div>	<div>31 Ga</div>	<div>32 Ge</div>	<div>33 As</div>	<div>34 Se</div>	<div>35 Br</div>	<div>36 Kr</div>	<div>37 Rb</div>	<div>38 Sr</div>	<div>39 Y</div>	<div>40 Zr</div>	<div>41 Nb</div>	<div>42 Mo</div>	<div>43 Tc</div>	<div>44 Ru</div>	<div>45 Rh</div>	<div>46 Pd</div>	<div>47 Ag</div>	<div>48 Cd</div>	<div>49 In</div>	<div>50 Sn</div>	<div>51 Sb</div>	<div>52 Te</div>	<div>53 I</div>	<div>54 Xe</div>	<div>55 Cs</div>	<div>56 Ba</div>	<div>57 La</div>	<div>58 Ce</div>	<div>59 Pr</div>	<div>60 Nd</div>	<div>61 Pm</div>	<div>62 Sm</div>	<div>63 Eu</div>	<div>64 Gd</div>	<div>65 Tb</div>	<div>66 Dy</div>	<div>67 Ho</div>	<div>68 Er</div>	<div>69 Tm</div>	<div>70 Yb</div>	<div>71 Lu</div>	<div>72 Hf</div>	<div>73 Ta</div>	<div>74 W</div>	<div>75 Re</div>	<div>76 Os</div>	<div>77 Ir</div>	<div>78 Pt</div>	<div>79 Au</div>	<div>80 Hg</div>	<div>81 Tl</div>	<div>82 Pb</div>	<div>83 Bi</div>	<div>84 Po</div>	<div>85 At</div>	<div>86 Rn</div>
<div>57 La</div>	<div>58 Ce</div>	<div>59 Pr</div>	<div>60 Nd</div>	<div>61 Pm</div>	<div>62 Sm</div>	<div>63 Eu</div>	<div>64 Gd</div>	<div>65 Tb</div>	<div>66 Dy</div>	<div>67 Ho</div>	<div>68 Er</div>	<div>69 Tm</div>	<div>70 Yb</div>	<div>71 Lu</div>	<div>72 Hf</div>	<div>73 Ta</div>	<div>74 W</div>	<div>75 Re</div>	<div>76 Os</div>	<div>77 Ir</div>	<div>78 Pt</div>	<div>79 Au</div>	<div>80 Hg</div>	<div>81 Tl</div>	<div>82 Pb</div>	<div>83 Bi</div>	<div>84 Po</div>	<div>85 At</div>	<div>86 Rn</div>																																																															
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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animals with similar looking features into families Starting in the 1990 s biologists have extracted DNA and RNA from cells as a guide to how plants or animals should be grouped Like visual structures these reveal the underlying design of creation Exploring the World of Biology is a fascinating look at life from the smallest proteins and spores to the complex life systems of humans and animals Chemistry is an amazing branch of science that affects us every day yet few people realize it or even give it much thought Without chemistry there would be nothing made of plastic there would be no rubber tires no tin cans no televisions no microwave ovens or something as simple as wax paper This book presents an exciting and intriguing tour through the realm of chemistry as each chapter unfolds with facts and stories about the discoveries of discoverers Find out why pure gold is not used for jewelry or coins Join Humphry Davy as he made many chemical discoveries and learn how they shortened his life See how people in the 1870s could jump over the top of the Washington Monument Exploring the World of Chemistry brings science to life and is a wonderful learning tool with many illustrations and biographical information

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