Reviewing Content

- The solvent is the substance in which the solute is dissolved.
- Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
- 44. solubility: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.

saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. unsaturated solution: a solution that contains less solute than a saturated solution at a given temperature and pressure. miscible: describes liquids that dissolve in each other. immiscible: describes liquids that are insoluble in each other.

- 45. Particles of solute crystallize.
- No: if there were undissolved solute, the excess solute would come out of a supersaturated solution.
- 47. 5.55 × 102 g AgNO₃
- 48. Solubility increases with pressure.
- 49. a. 1.6×10⁻² g/L b. 4.7×10⁻² g/L
- Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
- Molarity is the number of moles of solute dissolved in one liter of solution.
 - a. 1.3M KCI
 - b. 3.3 × 10⁻¹M MgCl₂
- 52. 2.00×10^{1} mL.
- a. 5.0×10⁻¹ mol NaCl, 29 g NaCl
 b. 1.0 mol KNO₃, 1.0×10⁻² g KNO₃
 c. 2.5×10² mol CaCl₂, 2.8 g CaCl₂
- 54. a. 2.3 × 10¹ g NaCl b. 2.0 g MgCl.
- 55. a. 16% (v/v) ethanol
 - b. 63.6% (v/v) isopropyl alcohol

- Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation. freezing-point depression, and vaporpressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution: relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
- 57. a. seawater
 - b. 1.5M KNO₃
 - e. 0.100M MgCl₂
- The effective molality of the Ca(NO₃)₂ solution is 3m. The effective molality of the NaNO₃ solution is 2m.
- 59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
- The salt lowers the freezing point of the ice-water cooling mixture.
- 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
- 62. Add 27.0 g H2O to 32.0 g CH2OH.
- 63. a. 100.26°C
 - b. 101.54°C
- 64. a. -4.46°C
 - b. -2.2°C
- 65. a. -1.1°C
 - b. -0.74°C
 - e. −1.5°C

Understanding Concepts

 a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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 - The Rise of Digital Reading World Of Chemistry Chapter 16 Test Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying World Of Chemistry Chapter 16 Test Answers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an World Of Chemistry Chapter 16 Test Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from World Of Chemistry Chapter 16 Test Answers
 - Personalized Recommendations
 - World Of Chemistry Chapter 16 Test Answers User Reviews and Ratings

- World Of Chemistry Chapter 16 Test Answers and Bestseller Lists
- 5. Accessing World Of Chemistry Chapter 16 Test Answers Free and Paid eBooks
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 - World Of Chemistry Chapter 16 Test Answers eBook Subscription Services
 - World Of Chemistry Chapter 16 Test Answers Budget-Friendly Options
- 6. Navigating World Of Chemistry Chapter 16 Test Answers eBook Formats
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 - Interactive Elements World Of Chemistry Chapter 16 Test Answers
- 8. Staying Engaged with World Of Chemistry Chapter 16 Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers World Of Chemistry Chapter 16 Test Answers
- 9. Balancing eBooks and Physical Books World Of Chemistry Chapter 16 Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection World Of Chemistry Chapter 16 Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine World Of Chemistry Chapter 16 Test Answers
 - Setting Reading Goals World Of Chemistry Chapter 16 Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of World Of Chemistry Chapter 16 Test Answers
 - Fact-Checking eBook Content of World Of Chemistry Chapter 16 Test Answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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