

Section 2: Stoichiometric Calculations - Notes

Objectives:

- List the sequence of steps used in solving stoichiometric problems.
- Solve stoichiometric problems.

Using Stoichiometry:

- **Steps to solve mole-to-mole, mole-to-mass, and mass-to-mass stoichiometric problems:**
 1. Write a balanced chemical equation for the reaction.
 2. Determine where to start your calculations by noting the unit of the given substance.
 - If mass (in grams) of the given substance is the starting unit, you must convert to moles.
 - If amount (in moles) of the given substance is the starting unit, convert moles of the given substance to moles of the unknown.
 3. The end point of the calculation depends on the desired unit of the unknown substance.
 - If the answer must be in moles, stop you are finished.
 - If the answer must be in grams, convert moles of unknown to grams of unknown using the molar mass as the conversion factor.
- **Example – Mole to Mole:** Sodium chloride forms by the reaction between sodium and chlorine. How many moles of NaCl result from the complete reaction of 3.4 mol of Cl_2 ?
 1. Write a balanced chemical equation:
 2. Decide which mole ratio to use:
 3. Use the mole ratio to convert the starting number of moles to final number of moles.

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Stoichiometric Calculations

American Ceramic Society



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