

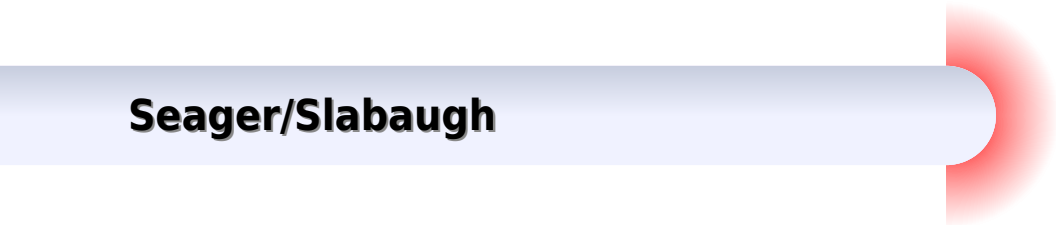
Ch 5 Ions and Ionic Compounds Study Guide**Multiple Choice**

Identify the choice that best completes the statement or answers the question.

- _____ 1. Which of the following electron configurations belongs to an element that is NOT chemically reactive?
- a. $1s^2 2s^2 2p^3$ c. $1s^2 2s^2 2p^5$
b. $1s^2 2s^2 2p^6 3s^2 3p^6$ d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- _____ 2. In the compound sodium fluoride, NaF, the sodium atom loses one electron and the fluorine atom gains one electron to form ions that have electron configurations similar to
- a. helium. c. neon.
b. oxygen. d. calcium.
- _____ 3. In many compounds, atoms of main-group elements form ions so that the number of electrons in the outermost energy levels of each ion is
- a. 2. c. 8.
b. 6. d. 10.
- _____ 4. The electron configuration of nitrogen is $1s^2 2s^2 2p^3$. How many more electrons does nitrogen need to have an electron configuration similar to neon?
- a. 1 c. 5
b. 3 d. 8
- _____ 5. Atoms of copper and iron
- a. generally form stable bonds with transition elements.
b. have stable electron configurations.
c. tend to form cations.
d. tend to form anions.
- _____ 6. An anion
- a. is an ion with a negative charge.
b. attracts ions with negative charges.
c. results when an alkaline-earth metal loses one of its two outermost electrons.
d. has more protons than electrons.
- _____ 7. When the octet rule is satisfied, the outermost _____ are filled.
- a. *d* and *f* orbitals c. *s* and *d* orbitals
b. *s* and *p* orbitals d. *d* and *p* orbitals
- _____ 8. The elements of the _____ group satisfy the octet rule without forming compounds.
- a. halide c. alkali metal
b. noble gas d. alkaline-earth metal
- _____ 9. Once an atom has full *s* and *p* orbitals in its outermost energy level,
- a. it is highly reactive only with alkali metals.
b. it is highly reactive only with halogens.
c. it can be combined with most elements.
d. it has a stable octet and is unreactive.

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