

25

NUCLEAR CHEMISTRY

Practice Problems

In your notebook, solve the following problems.

SECTION 25.1 NUCLEAR RADIATION

- What happens to the mass number and atomic number of an atom that undergoes beta decay?
- A nuclide of an element undergoes alpha particle decay. How do the atomic number and mass number of the particle change?
- Give the composition of the nucleus of the following isotopes.
 - $^{224}_{86}\text{Rn}$
 - $^{232}_{90}\text{Th}$
 - $^{209}_{83}\text{Bi}$
- Complete each of the following equations.
 - $^{210}_{82}\text{Pb} \rightarrow \text{?} + \text{?}$
 - $^{210}_{83}\text{Bi} \rightarrow \text{?} + \text{?}$
 - $^{238}_{92}\text{U} \rightarrow \text{?} + \text{?}$

SECTION 25.2 NUCLEAR TRANSFORMATIONS

- Write a nuclear equation for the following radioactive processes.
 - alpha decay of francium-223
 - electron capture by beryllium-7
 - beta emission by argon-37
 - positron emission by boron-12
- Complete the equations for these transmutation reactions.
 - $^{24}_{11}\text{Na} + \text{?} \rightarrow \text{?} + \text{?}$
 - $^{238}_{92}\text{U} + \text{?} \rightarrow \text{?} + \text{?} + \text{?}$
 - $^{240}_{94}\text{Pu} \rightarrow \text{?} + \text{?} + \text{?}$
 - $^{238}_{92}\text{U} \rightarrow \text{?} + \text{?} + \text{?} + \text{?}$
 - $\text{?} + \text{?} \rightarrow \text{?} + \text{?} + \text{?} + \text{?}$
- Po-214 has a relatively short half-life of 164 s. How many micrograms would it take for 0.5 g of this isotope to decay to 0.25 g?
- How many days does it take for 16 g of polonium-210 to decay to 1.0 g? The half-life of polonium-210 is 137 days.
- By approximately what factor would the mass of a sample of copper-66 decrease in 50 minutes? The half-life of copper-66 is 5.48 min.
- In 5.48 minutes, 1.00 g of argon-39 decays to leave only 0.12 g. What is the half-life of argon-39?

Study Guide Nuclear Radiation Answerd

**Katherine J. Denniston, Denniston
Katherine**



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Table of Contents Study Guide Nuclear Radiation Answerd

1. Understanding the eBook Study Guide Nuclear Radiation Answerd
 - The Rise of Digital Reading Study Guide Nuclear Radiation Answerd
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Nuclear Radiation Answerd
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Nuclear Radiation Answerd
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Nuclear Radiation Answerd
 - Personalized Recommendations
 - Study Guide Nuclear Radiation Answerd User Reviews and Ratings
 - Study Guide Nuclear Radiation Answerd and Bestseller Lists
5. Accessing Study Guide Nuclear Radiation Answerd Free and Paid eBooks
 - Study Guide Nuclear Radiation Answerd Public Domain eBooks
 - Study Guide Nuclear Radiation Answerd eBook Subscription Services
 - Study Guide Nuclear Radiation Answerd Budget-Friendly Options

6. Navigating Study Guide Nuclear Radiation Answerd eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Nuclear Radiation Answerd Compatibility with Devices
 - Study Guide Nuclear Radiation Answerd Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Nuclear Radiation Answerd
 - Highlighting and Note-Taking Study Guide Nuclear Radiation Answerd
 - Interactive Elements Study Guide Nuclear Radiation Answerd
8. Staying Engaged with Study Guide Nuclear Radiation Answerd
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Nuclear Radiation Answerd
9. Balancing eBooks and Physical Books Study Guide Nuclear Radiation Answerd
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Nuclear Radiation Answerd
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Nuclear Radiation Answerd
 - Setting Reading Goals Study Guide Nuclear Radiation Answerd
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Nuclear Radiation Answerd
 - Fact-Checking eBook Content of Study Guide Nuclear Radiation Answerd
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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