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**Bernadette Mazurek Melnyk, Ellen
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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160 0 \text{ g } \text{Al} \times 1 \text{ mol } \text{Al} / 27 0 \text{ g } \text{Al} \times 2 \text{ mol } \text{Al}_2\text{O}_3 / 4 \text{ mol } \text{Al} \times 101 964 \text{ g } \text{Al}_2\text{O}_3 / 2 \text{ mol } \text{Al}_2\text{O}_3 = 302 3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $260 0 \text{ g } \text{Al}_2\text{O}_3 / 302 3 \text{ g } \text{Al}_2\text{O}_3 \times 100 = 86 01$ theoretical 302 3 g

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 $80 \text{ g } \text{I}_2\text{O}_5 \cdot 1 \text{ mol } \text{I}_2\text{O}_5 / 1 \text{ mol } \text{I}_2 \cdot 253.8 \text{ g } \text{I}_2\text{O}_5 / 1 \text{ mol } \text{I}_2\text{O}_5 = 28 \text{ g } \text{CO} \cdot 1 \text{ mol } \text{CO} / 28 \text{ g } \text{CO} = 1 \text{ mol } \text{CO}$

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