

24. According to the balanced chemical equation,

How many moles of $\text{SO}_2(\text{g})$ will be produced when 1.5×10^7 molecules of zinc sulfide react with 1000 dm³ of oxygen gas? Assume a 75% yield.



25. I need to produce 500 g of lithium oxide(Li_2O)

- how many grams of lithium AHO
- how many liters of oxygen do I need

The balanced equation is: $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$

26. How many grams of water will be produced from 50 g hydrogen reacting with 50 g oxygen?

Think Critically:

22. The reaction of 1 mol of C to form carbon monoxide in the reaction $2 \text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2 \text{CO}(\text{g})$ releases 11.5 kJ of heat. How much heat will be released by the combustion of 100 g of C according to the above information?

23. According to the balanced chemical equation, how many atoms of silver will be produced from combining 300 g of copper with 200 g of silver nitrate?



24. According to the balanced chemical equation, how many moles of $\text{SO}_2(\text{g})$ will be produced when 1.5×10^7 molecules of zinc sulfide react with 1000 dm³ of oxygen gas? Assume a 75% yield.



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The balanced equation is: $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$

26. A tin ore contains 3.5% SnO_2 . How much tin is produced by reducing 2.0 kg of the ore with carbon?



27. If 96.5 g of HCl and 79 g of Zn are put together:



- Determine which reactant is the limiting reactant.
- Find the mass of ZnCl_2 formed.
- Find the volume of H_2 (g) at STP formed.
- Determine which reactant is in excess and by how much.

28. Many plants synthesize glucose by photosynthesis as follows:
 $\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) + \text{energy} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(\text{s}) + \text{O}_2(\text{g})$

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