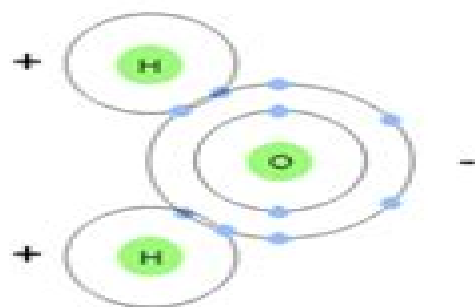


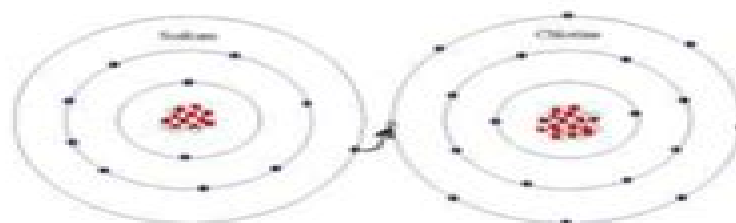
Chemical Bonding (PS 4c)

Test date: _____

Essential Concepts and Skills	Assignments
The Formation of Compounds through Ionic and Covalent Bonding (PS 4c) - Metals tend to lose electrons in chemical reactions, forming positive ions. Nonmetals tend to gain electrons in chemical reactions, forming negative ions. - Given a chemical formula of a compound, identify the elements and the number of atoms of each that comprise the compound. - Recognize that the number of electrons on the outermost energy level (valence electrons) determines an element's chemical properties or chemical reactivity. - Describe the difference between ionic and covalent bonding. - Predict what kind of bond (ionic or covalent) will likely form when metals and nonmetals are chemically bonded. - Gaining or losing electrons makes an atom an ion. - Atoms react to form chemically stable substances that are held together by chemical bonds and are represented by chemical formulas. To become chemically stable, atoms gain, lose, or share electrons. - Compounds are formed when elements chemically react. When a metallic element reacts with a nonmetallic element, their atoms gain and lose electrons respectively, forming ionic bonds. Generally, when two nonmetals react, atoms share electrons, forming covalent (molecular) bonds.	Pages 328-345 Starter Quiz 1, 2 Gizmos



Covalent Bond



Ionic Bond

Study Guide Ionic And Covalent Bonding

Roman Wölfel



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