

# Chemistry 1 Term 1 exam sample problems by topic

## physical or chemical change

boiling water

burning wood

melting aluminum

separating sand from water

P

P

P

## element names/symbols - know element symbols from handout during 1<sup>st</sup> week of class

## metric prefixes (centi-, milli-, kilo-)

Convert 54 mm to m.

Convert  $9.2 \times 10^{-5}$  cL to L.

Convert 5.12 kg to g.

$$\frac{54 \text{ mm}}{1000 \text{ mm}} = 0.054 \text{ m}$$

$$\frac{9.2 \times 10^{-5} \text{ cL}}{100 \text{ cL}} = 9.2 \times 10^{-7} \text{ L}$$

$$\frac{5.12 \text{ kg}}{1000 \text{ g}} = 5120 \text{ g}$$

## Temperature conversions

Convert 23 degrees C to Kelvin

Convert 50 K to degrees C

Convert 37 degrees C to degrees F

$$23 + 273 = 296 \text{ K}$$

$$50 - 273 = -223 \text{ K}$$

$$\frac{9}{5}(37 \text{ C}) + 32 = 98.6 \text{ F}$$

## significant digits

Round 8.9456 to 3 significant digits.

Round 513431 to 2 significant digits.

Add  $4.324 + 345.1$ , and round your answer properly.

Multiply  $32.56 \times 103$ , and round your answer properly.

$$8.95$$

$$510000$$

$$349.424 \rightarrow 349.4 \text{ (1 decimal place)}$$

$$3353.68 \rightarrow 3350 \text{ (3 sig figs)}$$

## scientific notation

Express 3712 in scientific notation.

Express 0.0377561 in scientific notation.

$$3.712 \times 10^3$$

$$3.77561 \times 10^{-2}$$

## density

What is the volume of a solid with a density of 7.8 g/mL and a mass of 11.8 g?

$$D = \frac{m}{V} \quad 7.8 \text{ g/mL} = \frac{11.8 \text{ g}}{V} \quad V = 1.51 \text{ mL}$$

## unit conversions (inches to cm, seconds to hours, etc) including imperial and metric

Convert 54 hours to seconds

Convert 84 oz to grams

Convert 1234 m to inches

Convert 2.3 quarts to Liters

$$\frac{54 \text{ hr}}{1 \text{ hr}} \times \frac{60 \text{ min}}{1 \text{ min}} \times \frac{60 \text{ sec}}{1 \text{ min}} = 194400 \text{ sec}$$

$$\frac{84 \text{ oz}}{16 \text{ oz}} \times \frac{1 \text{ lb}}{1 \text{ lb}} \times \frac{454 \text{ g}}{1 \text{ lb}} = 2400 \text{ g}$$

$$\frac{1234 \text{ m}}{1 \text{ m}} \times \frac{100 \text{ cm}}{1 \text{ m}} \times \frac{1 \text{ inch}}{2.54 \text{ cm}} = 48600 \text{ inches}$$

## Counting protons, neutrons, and electrons

| Symbol                       | Protons | Neutrons | Electrons | Charge |
|------------------------------|---------|----------|-----------|--------|
| ${}^4_2\text{Be}$            | 4       | 5        | 4         | 0      |
| ${}^{108}_{47}\text{Ag}$     | 47      | 61       | 47        | 0      |
| ${}^{201}_{80}\text{Hg}$     | 80      | 121      | 80        | 0      |
| ${}^{33}_{16}\text{S}^{2-}$  | 16      | 17       | 18        | -2     |
| ${}^{13}_{13}\text{Al}^{3+}$ | 13      | 15       | 10        | +3     |
| ${}^{86}_{34}\text{Se}^{-}$  | 34      | 47       | 35        | -1     |
| ${}^{60}_{28}\text{Ni}$      | 28      | 32       | 28        | 0      |
| ${}^{80}_{34}\text{Se}^{2-}$ | 34      | 46       | 36        | -2     |
| ${}^{16}_{8}\text{O}^{2-}$   | 8       | 8        | 10        | -2     |

$$\frac{2.3 \text{ qt}}{1.057 \text{ qt}} = 2.18 \text{ L}$$

# Study Guide Mixed Review Answer Key

**Patricia A. Potter, Anne G. Perry, Kyla  
C. Janzen**



## **Study Guide Mixed Review Answer Key:**

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## **Table of Contents Study Guide Mixed Review Answer Key**

1. Understanding the eBook Study Guide Mixed Review Answer Key
  - The Rise of Digital Reading Study Guide Mixed Review Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Mixed Review Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Mixed Review Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Mixed Review Answer Key
  - Personalized Recommendations
  - Study Guide Mixed Review Answer Key User Reviews and Ratings
  - Study Guide Mixed Review Answer Key and Bestseller Lists
5. Accessing Study Guide Mixed Review Answer Key Free and Paid eBooks
  - Study Guide Mixed Review Answer Key Public Domain eBooks
  - Study Guide Mixed Review Answer Key eBook Subscription Services
  - Study Guide Mixed Review Answer Key Budget-Friendly Options

6. Navigating Study Guide Mixed Review Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide Mixed Review Answer Key Compatibility with Devices
  - Study Guide Mixed Review Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide Mixed Review Answer Key
  - Highlighting and Note-Taking Study Guide Mixed Review Answer Key
  - Interactive Elements Study Guide Mixed Review Answer Key
8. Staying Engaged with Study Guide Mixed Review Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Mixed Review Answer Key
9. Balancing eBooks and Physical Books Study Guide Mixed Review Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Mixed Review Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Mixed Review Answer Key
  - Setting Reading Goals Study Guide Mixed Review Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Mixed Review Answer Key
  - Fact-Checking eBook Content of Study Guide Mixed Review Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

## **Study Guide Mixed Review Answer Key Introduction**

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