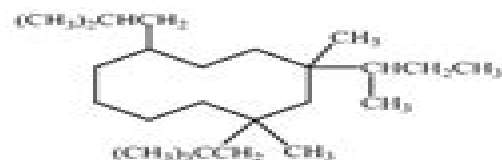


INSTRUCTIONS (Read carefully): Carefully separate all sheets. This examination has 2 pages of problems, 1 double-sided answer sheet, a periodic table, and 1 blank sheet for scratch work. Read each question carefully. For 2 **bonus points**, correctly complete as requested on *both sides* of the answer sheet: print your name, sign your name, give the portion of your FSU e-mail address that appears before the @ (this is the same as your BlackBoard login user name), give your row letter and seat number, and give your Discussion Section number. Put your answers in the spaces provided on the answer sheet. *All sheets will be collected; only answers on the answer sheets will be graded.* There is a total of 100 points plus 2 bonus points. Before the next lecture, go over problems about which you were not confident on this examination.

REGRADE POLICY: If you have a question about the number of points that you earned on your graded exam answer sheet, *do not make changes or additions to your returned, graded answer sheet.* Write on a *separate sheet of paper* what items you believe that you were not given sufficient credit and the reasons why you think so. Attach it to your original answer sheet and give it to Dr. Hilinski after a lecture or during an office hour, preferably within 2 weeks after it is returned. Be sure to make and keep a photocopy for your records.

- [9 points] For the compound given to the right, provide a name.
- [12 points] For each of the following, consider the compound given in *problem 1*.



- How many methylene hydrogens are there?
 - How many tertiary Cs are there?
 - Give the molecular formula of this compound.
 - For the *number of Cs* in this compound, what is the maximum number of Hs? Do not consider only the specific structural formula given in *problem 1*; do consider the general case for the number of Cs that you gave in *part 2c*.
 - How many tertiary Hs are there?
 - How many primary alkyl groups exist as substituents?
- [7 points] Provide a structural formula for 5-ethyl-2,3-dimethyl-8-(1-methylcyclopentyl)nonane. *For this problem, you must give a complete structural formula showing all atoms and all bonding valence electrons.*
 - [22 points] Consider the following.



(a) One of the 5 resonance structures (Lewis structures) that contribute to the resonance hybrid of the conjugate base of phenol is represented above as one of the products in the given chemical equation. Draw the additional 4 resonance structures (Lewis structures) that contribute to the resonance hybrid of the conjugate base of phenol.

Note that if you incorrectly duplicate resonance structures, you will lose credit for that resonance structure. If you give more than 4 resonance structures, points will be deducted.

For each of the four Lewis structures (resonance structures) that you give, you must show all atoms, all bonding valence electrons, all unshared valence electrons, and each nonzero formal charge.

(b) The equilibrium shown above lies to the right. This, of course, means that phenol is more acidic than cyclohexanol. Using structural considerations, provide a clear, concise explanation (but in as much detail as possible) why this equilibrium lies to the right. **NOTE** that a correct answer to this part is *not* that "the pK_a of phenol is less than the pK_a of cyclohexanol" because this is merely a restatement of the fact that phenol is more acidic than cyclohexanol.

- [10 points] Draw structural formulas for 5 (*and only 5*) constitutional isomers that *have one (and only one) ring* and the molecular formula C_4H_5FO . Yes, each of the constitutional isomers that you give must be *have one (and only one) ring*. Your structural formulas must have no nonzero formal charges and the most typical arrangements of bonding and nonbonding valence electrons. *Note that if you give the same isomer more than once, points will be deducted for that isomer. If you give more than 5 constitutional isomers, points will be deducted. For this problem, you must give complete structural formulas showing all atoms and bonding valence electrons. Do not give condensed structural formulas or bond-line formulas.*
- [4 points] A condensed structural formula for allene is shown to the right. Using lines, wedges, and dashed wedges for bonding electrons, give a 3-D stereoisomer for allene that is a representation of the true shape of this molecule. For your structure, label clearly each C with its hybridization (sp^3 , sp^2 , or sp).



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Ralph J. Fessenden, Joan S. Fessenden



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