

SHORT QUESTIONS

Q.1 Aromatic compounds burn with sooty flame, why?

Ans. Aromatic compounds have high percentage of carbon as compared to hydrogen. Aromatic compounds containing six carbon atoms have six hydrogen in it but the aliphatic compounds with six carbon atoms have fourteen hydrogen atoms aromatic compounds due to greater percentage of carbon, do not burn properly and give soot or smoke of carbon.

Q.2 What is meant by aromatic compounds?

Ans. The term aromatic was derived from the Greek word "aroma" meaning "fragrant" usually the compounds containing six membered benzene ring are called aromatic compound e.g., phenol, toluene, benzene, naphthalene, etc.

Q.3 Benzene usually undergoes electrophilic substitution reaction and not addition reactions, why?

Ans. In benzene ring, delocalized π -electrons are present in the form of sheath. This sheath of electrons above and below the carbon plane can be donated to the substance, which is electron deficient. During electrophilic substitution reactions, the aromaticity of ring remains as such. In addition reaction of benzene it changes to a non-aromatic compounds.

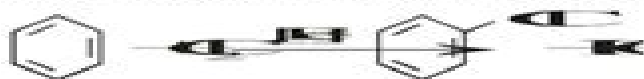


Q.4 Benzene is less reactive than alkanes, why?

Ans. In benzene ring, delocalized π -electrons are present in the form of sheath above and below the carbon plane. In case of ethene, localized π -electron cloud is present. Delocalized cloud of electrons is more stable as compared to localized cloud. Some energy is required to break the delocalization of electronic cloud called resonance energy. Due to this resonance stability benzene is less reactive than alkene.

Q.5 Give two reactions of benzene in which it behaves as saturated hydrocarbon.

Ans. (i) Benzene reacts with chlorine in the presence of Fe to form chlorobenzene.



Study Guide For Content Mastery Aromatic Hydrocarbons

Don Zimmerman



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