

My "Mole Concept Test" will take place on: \_\_\_\_\_!

*Before I write my "Mole Concept Test", I will be able to:*

- ☐ **explain why the mole is used in chemistry**
  - know how to find *atomic mass* on the periodic table
  - know the relationship between atomic mass and molar mass
  - recognize that a mole is a unit for counting atoms, ions, and molecules
- ☐ **perform calculations involving the mole**
  - determine the molar mass of an element or compound
  - convert among particles, moles, and mass
- ☐ **understand the relationship between moles and gases at STP**
  - know the volume of a mole of gas at STP (i. e. 22.4 L)
  - convert among moles, mass, and volume of a gas at STP
- ☐ **perform calculations involving molecular and empirical formulae**
  - distinguish between molecular and empirical formulae
  - determine the percent composition by mass from the formula of a compound
  - determine the empirical formula for a compound from its percent composition by mass
  - determine a molecular formula from the molecular mass and empirical formula
- ☐ **Describe concentration in terms of molarity**
  - Know that the units for molarity/concentration are mol/L or M
  - Write a procedure for how to prepare a solution of known molarity (a standard solution)
- ☐ **Perform calculations involving molarity**
  - Perform calculations relating mass, moles, volume, and molarity
- ☐ **Understand the meanings of the following vocabulary words:**
  - atomic mass
  - empirical formula
  - molarity
  - molar mass
  - molar solution
  - mole
  - molecular formula
  - molecular mass
  - quantitative transfer
  - percentage composition
  - standard solution
  - STP

# Study Guide Mole Test

**JA Banks**



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