

9.1

THE ARITHMETIC OF EQUATIONS

SECTION REVIEW

Objectives

- Calculate the amount of reactants required or product formed in a chemical process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

Key Terms

- stoichiometry

Part A Completion

Copy this completion exercise for class. Your understanding of the concepts and terms is clear when indicated in this exercise. Blank blanks are the completed work of others; absent phrases are none of these.

- The coefficients in a balanced chemical equation indicate _____ 1. _____ the relative number of _____ 2. _____ of reactants and products. All _____ 3. _____ stoichiometric calculations begin with a _____ 4. _____ On a _____ 5. _____ and _____ 6. _____ are conserved in every reaction; moles, volumes, and representative particles may not be.
- In solving stoichiometric problems, conversion factors relating moles of reactants to _____ 7. _____ of products are used.
- If you assume _____ 8. _____, the equation also tells you about the volume of gases.

Part B True-False

(Quickly mark off the statements as to be always true, AC; sometimes true, ST; or never true, NT.)

- _____ 7. The coefficients in a balanced chemical equation can be used to derive molar ratios relating reactants to products.
- _____ 8. The coefficients in a balanced chemical equation tell the relative volumes of reactants and products, expressed in any suitable units of volume.
- _____ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

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Arthur Wilberforce Jose



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