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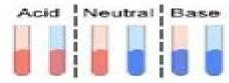
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Student Exploration: Titration

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: acid, analyte, base, dissociate, equivalence point, indicator, litmus paper, molarity, neutralize, pH, strong acid, strong base, titrant, titration, titration curve, weak acid, weak base

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)
There are several definitions of acids and bases. According to the Brensted-Lowry definition, an acid is a substance that is capable of donating a proton to another substance. A base is a substance that accepts protons. When an acid and a base are combined, the acid is neutralized as the base accepts the protons produced by the acid.



One way to determine if a solution is acidic or basic is to use **litmus paper**, as shown above. There are two types of litmus papers: red and blue.

How does litmus paper indicate an acid?

How does litmus paper indicate a neutral substance?

How does litmus paper indicate a base?

they turn red

one turns red the other turns blue

wellow.

ph is below 6

they turn blue

Gizmo Warm-up

Litmus is an example of an **indicator**, a substance that changes color depending on its **pH** (pH is a measure of the concentration of protons, or H* ions). In the Titration Gizmo, you will use indicators to show how acids are neutralized by bases, and vice versa.

To begin, check that 1.00 M NaOH is selected for the Burette, Mystery HBr is selected for the Flask, and Bromthymol blue is selected for the Indicator.

- 1. Look at the flask. What is the color of the bromthymol blue indicator?
- 2. What does this tell you about the pH of the solution in the flask?

Solutions with a pH below 7.0 are acidic, while those with a pH above 7.0 are basic

3. Move the slider on the burette to the top to add about 25 mL of NaOH to the flask. What happens, and what does this tell you about the pH of the flask?
it turns blue since its above 7.6

Titration Gizmo Answer Key Teacher Guide

Joseph M. Oxenhorn, Michael Norman Idelson, Peter Greenleaf

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