

Name: \_\_\_\_\_

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**Titration Practice Worksheet**

*Find the requested quantities in the following problems:*

- 1) Write a balanced equation for the reaction between NaOH and HCl. Use this equation to answer problems 2-4
  
- 2) If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?
  
- 3) If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?
  
- 4) A 25.0 mL sample of HCl was titrated to the endpoint with 15.0 mL of 2.0 M NaOH. What is the molarity of HCl?
  
- 5) Write a balanced equation for the reaction between  $\text{Ca(OH)}_2$  and  $\text{HNO}_3$ . Use this equation to answer problems 6 -8.
  
- 6) If it takes 50.0 mL of 0.5 M  $\text{Ca(OH)}_2$  solution to completely neutralize 125 mL of  $\text{HNO}_3$  solution, what is the concentration of the  $\text{HNO}_3$  solution?
  
- 7) How many mL of 0.50 M  $\text{HNO}_3$  is necessary to titrate 25.0 mL of 0.05 M  $\text{Ca(OH)}_2$  solution to the endpoint?
  
- 8) If it takes 75.0 mL of 1.5 M  $\text{HNO}_3$  solution to completely neutralize 125 mL of  $\text{Ca(OH)}_2$ , what is the concentration of the  $\text{Ca(OH)}_2$  solution?

# Titration Worksheet Chemfiesta

**E. P. Serjeant**



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