Electron	Configuration	Practice
Chemistr	7	

Nome	:		
Due D	ate:		

How to write an electron configuration:

- A. Determine the total number of electrons to be represented.
- B. Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- C. The sum of the superscripts should equal the total number of electron. Example: $12Mg\ ls^2\ 2s^2\ 2p^6\ 3s^2$

I. Configuration Writing Practice

- A. Write a ground state electron configuration for each neutral at: 1. Ground state means that all of the lowest possible energy levels (up to the pr per number of electrons for the element) are filled.
- 1. No: 1522522p6 351
- 2. Pb: 1522522p63523p64523d104p65524d105p66524f145d106p2
- 3. 5r: 1522522p63523p64523d104p6552
- 4. U: 1522522p63523p64523d104p65524d105p66524f145d106p67525f36d1
- 5. N: 152 252 2p3
- *6. Ag: 1522522p63523p64523d104p65514d10
 - 7. Ti: 1522522p6 3523p6 452 3d2
 - 8. ce: 1522522p63623p64523d104p65524d105p66524f15d1
 - 9. CI: 1522522p63523p5
 - 10. Hg: 1522522p6 3523p6 4523d104p6 5524d105p6 6524f145d10

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Holt Chemistry Ralph Thomas Myers, 2004

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