

#### **D - Molar Mass**

79. Consider the following:

a. If the formula of a compound is known (e.g.,  $C_4H_4OCl$ ), and you find the masses of the component atoms from the Periodic Table ( $C = 12.011g$ ,  $H = 1.0079$ ,  $O = 15.9994$ ,  $Cl = 35.453g$ ), HOW would you find the molar mass of the compound?

b. Calculate the molar mass of one mole of  $C_4H_4OCl$ .

c. Calculate the mass of 5 moles of  $C_4H_4OCl$ .

#### **E - Percent Composition of Compounds**

80. Provide the word equation describing how you solve the mass percent for a given element in a compound.

90. Penicillin is an important antibiotic that has the chemical formula  $C_{16}H_{18}N_2SO_4$ . Calculate the mass percent of EACH element in this compound.

#### **F - Empirical Formulas**

100. Consider the following:

a. Taking in to account the formula for glucose,  $C_6H_{12}O_6$ , what information can be gathered about this compound?

b. What ratio of  $C:H:O$  is present when considering the formula above?

c. Now take in to account the following formulas:  $C_4H_8O_4$ ,  $C_2H_4O_2$ , and  $C_3H_6O_3$ . What do these formulas have in common with glucose?

113. Using the word "ratio" in the descriptions, explain what the difference is between empirical and

# Skill Practice 31 Moles And Formulas Answer Key

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Gislason, Sheila McNicholas**



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