Name: \_\_\_\_\_\_ Date: \_\_\_\_\_\_

## **Net Ionic Equations: Practice Problems**

Balance the following equations. Show the ionic and net ionic forms.

- Mg(NO<sub>3</sub>)<sub>2</sub>(aq.) + Na<sub>2</sub>CrO<sub>4</sub>(aq.) → MgCrO<sub>4</sub>(s) + NaNO<sub>3</sub>(aq.) lonic:
   Net ionic:
- 2) Na<sub>3</sub>PO<sub>4</sub>(aq.) + Ni(ClO<sub>4</sub>)<sub>2</sub>(s) → Ni<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>(s) + NaClO<sub>4</sub>(aq.) lonic:
  Net ionic:
- 3) AgNO<sub>3</sub>(aq.) + MgI<sub>2</sub>(aq.) → AgI(s) + Mg(NO<sub>3</sub>)<sub>2</sub>(aq.) lonic:
  Net ionic:
- 4) Ca(OH)₂(aq.) + FeCl₃(aq.) → Fe(OH)₃(s) + CaCl₂(aq.) lonic:
  Net ionic:
- 5) RbF(aq.) + CuSO₄(aq.) → Rb₂SO₄(aq.) + CuF₂(aq.) lonic:
  Net ionic:
- 6) (NH<sub>4</sub>)<sub>2</sub>S(aq.) + Co(NO<sub>3</sub>)<sub>2</sub>(aq.) → CoS(s) + NH<sub>4</sub>NO<sub>3</sub>(aq.) lonic:
  Net ionic:
- 7) Na<sub>3</sub>PO<sub>4</sub>(aq.) + FeCl<sub>3</sub>(aq.) → FePO<sub>4</sub>(s) + NaCl(aq.) lonic:
  Net ionic:
- 8) AgNO<sub>3</sub>(aq.) + HCI(aq.) → AgCI(s) + HNO<sub>3</sub>(aq.) lonic:
  Net ionic:
- 9) MgBr₂(aq.) + KOH(aq.) → Mg(OH)₂(s) + KBr(aq.) lonic: Net ionic:
- 10) AlCl<sub>3</sub>(aq.) + K<sub>3</sub>PO<sub>4</sub>(aq.) → AlPO<sub>4</sub>(s) + KCl(aq.) lonic:
  Net ionic:

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