

1.2 Mole Conversions Solutions

PRACTICE EXAMPLES- To complete the practice examples (1) write the type of question you have, (2) write the equation with units, (3) show the cancellation of units and (4) calculate the solution.

1. Determine how many atoms are present in 2.5 moles of Silicon? (**15.05×10^{23} atoms**)

MOLES $\rightarrow$ ATOMS

$$\frac{2.5 \text{ moles}}{1} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 15.05 \times 10^{23} \text{ atoms}$$

2. How many moles of Silicon is 3.01×10^{23} atoms? (**.5 moles**)

ATOMS $\rightarrow$ MOLES

$$\frac{3.011 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ mole}}{6.022 \times 10^{23} \text{ atom}} = .5 \text{ moles}$$

3. What is the mass of 4 moles of Iron? (**224 grams**)

MOLES $\rightarrow$ GRAMS

$$\frac{4 \text{ moles}}{1} \times \frac{55.9 \text{ grams}}{\text{mole}} = 224 \text{ grams of C}$$

4. How many atoms are in 48.6 grams of Magnesium? (**12.04×10^{23} atoms**)

GRAMS $\rightarrow$ ATOMS

$$\frac{48.6 \text{ grams}}{1} \times \frac{1 \text{ mole}}{24.3 \text{ grams}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} \approx 12.04 \times 10^{23} \text{ atoms}$$

5. How many grams are in 18.06×10^{23} atoms of Copper? (**190.5 grams**)

ATOMS $\rightarrow$ GRAMS

$$\frac{18.06 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ mole}}{6.022 \times 10^{23} \text{ atoms}} \times \frac{63.5 \text{ grams}}{1 \text{ mole}} \approx 190.5 \text{ grams}$$

6. How many moles of Nickel is 176 grams? (**3 moles**)

GRAMS $\rightarrow$ MOLES

$$\frac{176 \text{ grams of C}}{1} \times \frac{\text{mole}}{58.7 \text{ grams}} \approx 3 \text{ moles}$$

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A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the 'EduGorilla Prep Experts' text.

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Table of Contents Skill Practice 40 Moles And Formulas Practice

1. Understanding the eBook Skill Practice 40 Moles And Formulas Practice
 - The Rise of Digital Reading Skill Practice 40 Moles And Formulas Practice
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 40 Moles And Formulas Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 40 Moles And Formulas Practice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 40 Moles And Formulas Practice
 - Personalized Recommendations
 - Skill Practice 40 Moles And Formulas Practice User Reviews and Ratings
 - Skill Practice 40 Moles And Formulas Practice and Bestseller Lists
5. Accessing Skill Practice 40 Moles And Formulas Practice Free and Paid eBooks
 - Skill Practice 40 Moles And Formulas Practice Public Domain eBooks
 - Skill Practice 40 Moles And Formulas Practice eBook Subscription Services
 - Skill Practice 40 Moles And Formulas Practice Budget-Friendly Options
6. Navigating Skill Practice 40 Moles And Formulas Practice eBook Formats
 - ePub, PDF, MOBI, and More
 - Skill Practice 40 Moles And Formulas Practice Compatibility with Devices
 - Skill Practice 40 Moles And Formulas Practice Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 40 Moles And Formulas Practice
 - Highlighting and Note-Taking Skill Practice 40 Moles And Formulas Practice
 - Interactive Elements Skill Practice 40 Moles And Formulas Practice
8. Staying Engaged with Skill Practice 40 Moles And Formulas Practice

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Skill Practice 40 Moles And Formulas Practice
- 9. Balancing eBooks and Physical Books Skill Practice 40 Moles And Formulas Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 40 Moles And Formulas Practice
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 40 Moles And Formulas Practice
 - Setting Reading Goals Skill Practice 40 Moles And Formulas Practice
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 40 Moles And Formulas Practice
 - Fact-Checking eBook Content of Skill Practice 40 Moles And Formulas Practice
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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