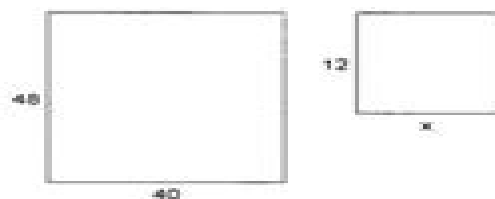


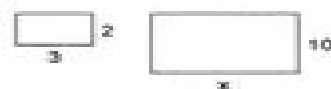
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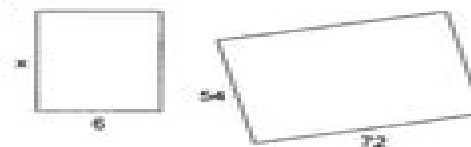
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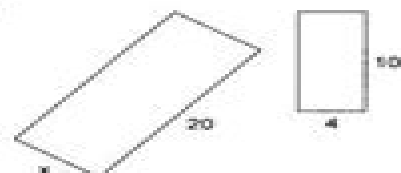
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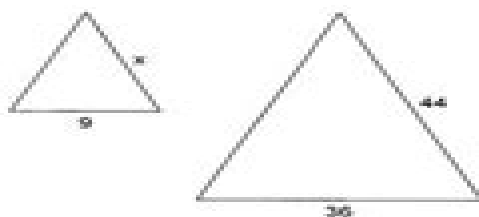
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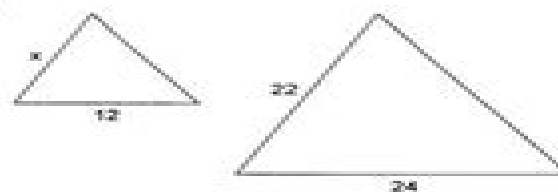
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web the substances that undergo change are called reactants the new substances are products sometimes during a chemical reaction one type of reactant will be used up before the other reactants this reactant is the limiting reactant using the limiting reactants gizmo you can determine which reactant is limiting in various scenarios

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web the limiting reagent concept allows us to calculate amounts of reactants used and products formed in a complete chemical reaction based on the stoichiometric relationships in a balanced chemical equation so far we have considered these relationships on the basis of masses and moles

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web 33mol o₂ 6mol br₂ 11mol o₂ 18mol br₂ identify the limiting reactant s and excess reactant s the limiting reactant is o₂ since it would yield the least amount of product 18 mol br₂ the excess reactant is c 2 h 3 br 3 since its complete reaction would have yielded up to 54 mol br₂ think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o₂ 0 2 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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web if you start with 14 82 g of ca oh 2 c a o h 2 and 16 35 g of h₂so₄ h 2 s o 4 a determine the limiting reagent b determine the number of moles of h₂o h 2 o produced c determine the number of grams of caso₄ c a s o 4 produced d determine the

number of grams of excess reagent left 1 make sure the equation is balanced

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