Name:

Use the periodic table to find the two ions (positive ion and negative ion) for each compound. Then write the formula.

Remember, the overall charge of the compound is neutral (zero charge). Thus, the overall positive charge must be equal to the overall negative charge.

Example: calcium chloride

 Ca^{2+} Cl^{-} (These two charges are not the same, but 2 chloride ions will give an overall negative charge of -2. The positive ion already has a + 2 charge.)

Ca2+Cl2

Ca²*

Count up the total charge. +2, -1, -1 This adds up to zero!!

Therefore, CaCl2 is the answer.

Name	Positive Ion	Negative Ion	Formula
1) sodium chloride ex.	Na *	ci-	NaCl
2) potassium bromide	-		
3) calcium oxide			
4) magnesium sulfide			
5) zinc chloride			
6) sodium sulfide	-		
7) strontium fluoride	1.7		
8) lithium oxide			
9) potassium nitride			
10) sodium phosphide			
II) aluminum chloride			
(2) calcium nitride			
13) aluminum oxide		T T	
(4) magnesium phosphide			
l5) aluminum sulfide			

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Table of Contents Skills Worksheet Math Skills Writing Ionic Formulas Answers

- 1. Understanding the eBook Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - The Rise of Digital Reading Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Personalized Recommendations
 - Skills Worksheet Math Skills Writing Ionic Formulas Answers User Reviews and Ratings
 - Skills Worksheet Math Skills Writing Ionic Formulas Answers and Bestseller Lists

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 - Highlighting and Note-Taking Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Interactive Elements Skills Worksheet Math Skills Writing Ionic Formulas Answers
- 8. Staying Engaged with Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skills Worksheet Math Skills Writing Ionic Formulas Answers
- 9. Balancing eBooks and Physical Books Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skills Worksheet Math Skills Writing Ionic Formulas Answers
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Setting Reading Goals Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Fact-Checking eBook Content of Skills Worksheet Math Skills Writing Ionic Formulas Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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