

The equilibrium constant, K , for a redox reaction is related to the standard potential, E° , by the equation

$$\ln K = \frac{nFE^\circ}{RT}$$

where n is the number of moles of electrons transferred, F (the Faraday constant) is equal to $96,500 \text{ C}/(\text{mol } e^-)$, R (the gas constant) is equal to $8.314 \text{ J}/(\text{mol} \cdot \text{K})$, and T is the Kelvin temperature.

11

Use the table of standard reduction potentials given above to calculate the equilibrium constant at standard temperature (25 °C) for the following reaction:



Express your answer numerically.

$K =$

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1998

Calculate the standard cell potential (E°) for the reaction



1000

Express your answer to three significant figures and include the appropriate units.

$E =$

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