Exercise No. 1: Alkali and Alkaline-earth Elements

Introduction

It is characteristic of the alkali elements that their cations do not form insoluble chlorides, sulfides, carbonates, or phosphates. Therefore, in the qualitative analysis scheme, they do not precipitate when successive reagents are added but are left in the residual solution as unprecipitated ions. They are usually identified by their flame spectra in the presence of HCl as a volatilizing agent. Because the flame test is made after separation into all the groups and the test is sensitive to the presence of even trace amounts of elements, it is necessary to guard against the introduction of contaminating amounts of alkali element when adding the precipitating reagents. For example, not only is it inadmissible to add NaCl to precipitate the insoluble chlorides but the HCl used instead must be checked for possible presence of Na contamination. In the present scheme, only the most common alkali elements (Na and K) are included.

As stated, cations of the alkaline-earth elements do not form insoluble chlorides or sulfides. Except for beryllium, which is not included in the present scheme, and magnesium, they do form insoluble carbonates and can thus be separated from the alkali elements. Once the carbonates have been precipitated, they can be dissolved again in an acetic acid buffer and the alkaline-earth elements precipitated out consecutively, making use of differences in the solubilities of the chromates and oxalates. Because the difference with strontium is so slight, it is not included in the present scheme.

Learning Outcomes

 Differentiate alkali from alkaline – earth elements through flame tests and solubility.

Materials

- Laboratory Manual
- Bunsen Burner
- Hotplate and water bath
- Chemicals and Glasswares:
 - a. 50 mL beaker
 - b. Test tubes
 - Medicine droppers

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Marcel A. Müller

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