

Practice Problems: Quantum Theory

CHEM 1A

1. Hydrogen has a red emission line at 656.3 nm, what is the energy and frequency of a photon of this light? Note: Planck's constant = 6.626×10^{-34} J·s, speed of light = 2.998×10^8 m/s

$$E_{\text{photon}} = h\nu = \frac{hc}{\lambda}$$

$$E_{\text{photon}} = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s}) (2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$E_{\text{photon}} = 3.02677861 \times 10^{-19} \text{ J}$$

$$\nu = \frac{c}{\lambda}$$

$$\nu = \frac{(2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$\nu = 4.568032912 \times 10^{14} \text{ s}^{-1}$$

Answer (E): $3.027 \times 10^{-19} \text{ J}$ Answer (ν): $4.568 \times 10^{14} \text{ s}^{-1}$ or Hz

2. What is the de Broglie wavelength for an electron moving at 2.42×10^6 m/s?
Note: $m_e = 9.109 \times 10^{-31}$ kg

$$\lambda = \left(\frac{h}{mv} \right)$$

$$\lambda = \left(\frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(9.109 \times 10^{-31} \text{ kg}) (2.42 \times 10^6 \text{ m/s})} \right) = 3.00583657 \times 10^{-10} \text{ m}$$

Answer: $3.01 \times 10^{-10} \text{ m}$ \longleftarrow many orders of magnitude larger than the diameter of an electron

3. Using the Bohr equation, calculate is the change in energy when an electron transitions from $n = 3$ to $n = 2$ in a hydrogen atom.

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left(\frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{initial}}^2} \right)$$

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left(\frac{1^2}{2^2} - \frac{1^2}{3^2} \right)$$

$$\Delta E = -3.02777778 \times 10^{-19} \text{ J}$$

Answer: $-3.03 \times 10^{-19} \text{ J}$

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The Princeton Review



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Table of Contents Study Guide Quantum Theory Answers

1. Understanding the eBook Study Guide Quantum Theory Answers
 - The Rise of Digital Reading Study Guide Quantum Theory Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Quantum Theory Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Quantum Theory Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Quantum Theory Answers
 - Personalized Recommendations
 - Study Guide Quantum Theory Answers User Reviews and Ratings
 - Study Guide Quantum Theory Answers and Bestseller Lists
5. Accessing Study Guide Quantum Theory Answers Free and Paid eBooks
 - Study Guide Quantum Theory Answers Public Domain eBooks
 - Study Guide Quantum Theory Answers eBook Subscription Services
 - Study Guide Quantum Theory Answers Budget-Friendly Options

6. Navigating Study Guide Quantum Theory Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Quantum Theory Answers Compatibility with Devices
 - Study Guide Quantum Theory Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Quantum Theory Answers
 - Highlighting and Note-Taking Study Guide Quantum Theory Answers
 - Interactive Elements Study Guide Quantum Theory Answers
8. Staying Engaged with Study Guide Quantum Theory Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Quantum Theory Answers
9. Balancing eBooks and Physical Books Study Guide Quantum Theory Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Quantum Theory Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Quantum Theory Answers
 - Setting Reading Goals Study Guide Quantum Theory Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Quantum Theory Answers
 - Fact-Checking eBook Content of Study Guide Quantum Theory Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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