LUTION CONCENTRATION - MASS X: Mass solute +100 -PPT: Equation above x 1000 -PPM: X1,000,000 -PPB: X 1,000,000,000 -MOLE FRACTION: X= moles solution - MOLARITY: M- mol solute L SOWHEN - MOLALITY: m = moles solute **VAPOR PRESSURE OF A** - NONVOLATILE = NO VP - ADDING NONVOLATILE SOLUTE TO SOLVENT LOWERS VP B/C IT TAKES UP SPACE IN THE SOLUTION & INTERFERS W/ ESCAPE OF SOLVENT - RAOULT'S LAW SHOWS RELATIONSHIP B/W SOLVENT VP & CONCENTRATION solvent DUCK control

FACTORS AFFECTING SOLUBILITY

- "LIKE DISSOLVES LIKE"
- POLAR SOLVENTS TEND TO DISSOLVE POLAR MOLECULES &
 - DISSOLVE POLAR MOLECULES & IONIC CMPDS
 - B/C POS/NEG END OF POLAR SOLVENT (LIKE H2O) ATTRACTED TO POS/NEG END
- OF POLAR MOLECULE
 -NONPOLAR SOLVENTS DISSOLVE
 NONPOLAR MOLECULES
- -FOR SOLIDS, SOLUBILITY AS
 TEMP ★
- FOR GAS, SOLUBILITY AS TEMPA
- FOR LIQUID AND SOLID: PRESSURE HAS LITTLE AFFECT ON SOLUBILITY IN H2O
- FOR GAS: HIGHER PARTIAL PRESSURE, THE MORE SOLUBLE -DEFINE BY HENRY'S LAW:

VAPOR PRESSURE OF SOLUTION W/ VOLATILE - IN IDEAL SOLUTION WHERE

- SOLUTE AND SOLVENT FOLLOW RAOULT'S, VP=SUM OF VP FROM
- DO RAQULT'S LAW TWICE:
- PB= XB . PBO & PA+P

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Dr Rajesh Thakur, Dr S R Singh, Subhash Jain, Mamta Mehrotra

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