

**Solutions** are homogeneous mixtures of two or more substances. The substance present in smaller amount in the mixture is called the **solute** while the substance that has greater amount is the **solvent**. The solute is dissolved in the solvent. Homogenous means that you cannot identify the particles of the solute and the solvent separately in the solution. Mixture means that the solute and the solvent retain their individual properties and that the solute and the solvent are only physically combined. A solution can be gaseous like air, liquid like soda or solid like metal brass (alloy of copper and zinc).

**Aqueous solutions** are solutions in which the solvent is water.

**Electrolytic solutions** are solutions that can conduct electricity due to the presence of dissolved ions.

**Non-electrolytic solutions** are solutions that cannot conduct electricity because the solutes do not ions when dissolved in water.

**Electrolyte** is a solute that results to an electrolytic solution when dissolved in water.

**Non-electrolyte** is a solute that does not result to an electrolytic solution when dissolved in water.

In strong electrolytic solutions (solutions of strong electrolytes), the electrolyte is assumed to be 100% dissociated. **Dissociation** pertains to the separation of positive and negative ions in an electrolyte. If we have a strong electrolyte like Potassium Nitrate ( $\text{KNO}_3$ ), what this means is that  $\text{KNO}_3$  completely separates to  $\text{K}^+$  and  $\text{NO}_3^-$  when dissolved in water. No  $\text{KNO}_3$  is left undissociated.

The table below shows strong and weak electrolytes as well as non-electrolytes.

Strong electrolytes	Weak electrolytes	Non-electrolytes
HCl	$\text{CH}_3\text{COOH}$	$(\text{NH}_2)_2\text{CO}$
$\text{HNO}_3$	HF	$\text{CH}_3\text{OH}$
$\text{HClO}_4$	$\text{HNO}_2$	$\text{C}_2\text{H}_5\text{OH}$
$\text{H}_2\text{SO}_4$	$\text{NH}_3$	$\text{C}_6\text{H}_{12}\text{O}_6$
NaOH	$\text{H}_2\text{O}$	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$
$\text{Ba}(\text{OH})_2$		
Ionic compounds		

Therefore, NaOH completely dissociates to form  $\text{Na}^+$  and  $\text{OH}^-$



The (aq) subscript indicates that the ions are dissolved in water. Notice that the arrow is unidirectional which means that the reverse reaction does not happen spontaneously.

For weak electrolytes like HF, the process is reversible and eventually reaches an equilibrium state in which the rate with which HF dissociate to  $\text{H}^+$  and  $\text{F}^-$  ions is equal to the rate with which  $\text{H}^+$  and  $\text{F}^-$  ions recombine to form HF.



### Precipitation Reactions

Precipitation reactions are reactions that lead to the formation of precipitates which are water insoluble, solid substances.

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## **Table of Contents Study Guide Reactions In Aqueous Solutions Answers**

1. Understanding the eBook Study Guide Reactions In Aqueous Solutions Answers
  - The Rise of Digital Reading Study Guide Reactions In Aqueous Solutions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Reactions In Aqueous Solutions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Reactions In Aqueous Solutions Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Reactions In Aqueous Solutions Answers
  - Personalized Recommendations
  - Study Guide Reactions In Aqueous Solutions Answers User Reviews and Ratings
  - Study Guide Reactions In Aqueous Solutions Answers and Bestseller Lists
5. Accessing Study Guide Reactions In Aqueous Solutions Answers Free and Paid eBooks
  - Study Guide Reactions In Aqueous Solutions Answers Public Domain eBooks
  - Study Guide Reactions In Aqueous Solutions Answers eBook Subscription Services
  - Study Guide Reactions In Aqueous Solutions Answers Budget-Friendly Options
6. Navigating Study Guide Reactions In Aqueous Solutions Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide Reactions In Aqueous Solutions Answers Compatibility with Devices
  - Study Guide Reactions In Aqueous Solutions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide Reactions In Aqueous Solutions Answers
  - Highlighting and Note-Taking Study Guide Reactions In Aqueous Solutions Answers
  - Interactive Elements Study Guide Reactions In Aqueous Solutions Answers
8. Staying Engaged with Study Guide Reactions In Aqueous Solutions Answers



- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Reactions In Aqueous Solutions Answers
- 9. Balancing eBooks and Physical Books Study Guide Reactions In Aqueous Solutions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Reactions In Aqueous Solutions Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Reactions In Aqueous Solutions Answers
  - Setting Reading Goals Study Guide Reactions In Aqueous Solutions Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Reactions In Aqueous Solutions Answers
  - Fact-Checking eBook Content of Study Guide Reactions In Aqueous Solutions Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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