

Stoichiometry Study Guide KEY

Chemistry

RHS – Mr. Moss

1. Define the following:
 - a. *Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.*
 - b. *Mole –The SI unit used to measure the amount of a substance that contains 6.02×10^{23} atoms of that substance.*
 - c. *Mole Ratio-The ratio between any two substances in a balanced chemical equation.*
 - d. *Molar Mass-The mass in grams of one mole of any substance.*
2. What relationships can be determined from a balanced chemical equation? *The relationships between particles, moles, and mass for all reactants and products.*
3. Explain why mole ratios are central to stoichiometric calculations. *Mole ratios allow for the conversion from moles of one substance in a balanced chemical equation to moles of another substance in the same equation.*
4. What is the first step in all stoichiometric calculation? (Be careful.) *Write a balanced chemical equation.*
5. How is molar mass used in some stoichiometric calculations? *Molar mass is a conversion factor for converting moles of a given substance to mass or mass of a given substance to moles.*
6. Find the molar mass of the following:
 - a. H_2SO_4 (sulfuric acid) *98.09 g/mol*
 - b. ZnO (zinc oxide) *81.39 g/mol*
 - c. NH_4NO_3 (ammonium nitrate) *80.05 g/mol*
7. List the mole ratios of the following *balanced* chemical equation.
 - a. $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$ $\frac{2 \text{ mol Al}}{3 \text{ mol Br}_2} ; \frac{3 \text{ mol Br}_2}{2 \text{ mol AlBr}_3} ; \frac{2 \text{ mol Al}}{2 \text{ mol AlBr}_3}$ *(Each one of these can be inverted to form another mole ratio. This would give a total of 6 ratios.)*

Study Guide Stoichiometry Key

James R. Braun, Sherman



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General, Organic, and Biochemistry Study Guide Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian s Organic and Biochemistry Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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