

**Model Question Paper-1 with effect from 2018-19
(CBCS Scheme)**

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18CHE12/22

**First/Second Semester B.E. Degree Examination
Engineering Chemistry**

(Common to all Branches)

Time : 3 Hrs

Max. Marks: 100

Note: Answer any FIVE full questions, choosing one full question from each module

Module-1

1. a. Define the terms: (i) Free energy, (ii) Entropy & (iii) Cell potential (6 Marks)
- b. What are concentration cells? The cell potential of Cu concentration cell $\text{Cu}/\text{CuSO}_4 (0.005\text{M}) // \text{CuSO}_4 (X) / \text{Cu}$ is 0.0295 V at 25° C. Write the cell reaction and calculate the value of X. (7 Marks)
- c. What are batteries? Explain the construction and working of Li-ion battery, mention its applications. (7 Marks)

OR

2. a. Write short notes on primary, secondary and reserve batteries. (6 Marks)
- b. Explain the construction and working of Ni-MH battery. Mention its applications. (7 Marks)
- c. For the cell, $\text{Fe}/\text{Fe}^{2+} (0.01\text{M}) // \text{Ag}^+ (0.1\text{M}) / \text{Ag}$ write the cell reaction and calculate the emf of the cell at 298K, if standard electrode potentials of Fe and Ag electrodes are -0.44V and 0.8 V respectively. (7 Marks)

Module-2

3. a. What is metallic corrosion? Describe the electrochemical theory of corrosion taking iron as an example. (7 Marks)
- b. Explain: (i) Differential metal corrosion & (ii) Water-line corrosion (6 Marks)
- c. What is electroplating? Explain the electroplating of chromium (7 Marks)

OR

4. a. What is meant by metal finishing? Mention (any five) technological importance of metal finishing. (6 Marks)
- b. What is electrolessplating? Explain the electrolessplating of copper. (8 Marks)
- c. Explain the factors affecting the rate of corrosion (i) Nature of corrosion product, (ii) Ratio of anodic to cathodic areas & (iii) pH (6 Marks)

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