

CP6 Radioactivity

Background Radiation

What is natural background radiation?

State 2 sources of natural background radiation:

Draw a labelled diagram of Rutherford's model State 2 sources of background radiation that does not occur naturally

How can radiation be measured and describe how the equipment works.

Describe the experiment Rutherford carried out which allowed him to come up with his model.

Types of radiation

Describe the structure of an alpha particle What is the radius of a nucleus?

What is the radius of an atom?

What is the charge on an alpha particle?

How is Rutherford's model of the atom similar to Bohr's model? Describe the structure of a beta particle

What is the charge on a beta particle?

Inside Atoms

Draw a labelled diagram of the modern day structure of an atom Describe the structure of a gamma ray

What does unstable mean?

	Mass	Charge
Proton		
Neutron		
Electron	Alpha	Beta
Gamma		Weak

Complete the table

Penetration	C	N	Mg	Co
Mass number				
Atomic number				
Protons				
Neutrons				

Explain why alpha particles are more ionising than beta particles

Define an isotope

Radioactive decay

Complete the table

Particle	Symbol
Alpha	
Beta	β^-
Positron	
Neutron	n

Describe how a magnesium atom becomes a Mg^{2+} ion. What happens to the atomic and mass number when a neutron is emitted?

Electrons and Orbits

Radioactive decay happens when an atom absorbs electromagnetic radiation Describe what happens during β^- decay and the effect on the nucleus

Describe what happens when an atom emits electromagnetic radiation Describe what happens during β^- decay and the effect on the nucleus

What is the difference between emission and absorption spectra Complete the table summarising what happens to the mass and atomic number during radioactive decay

	Mass number	Atomic number
Alpha		
Beta		Increases by 1

Radium has a mass number of 226 and an atomic number of 88.

Write a nuclear equation when this nucleus emits an alpha particle

Write a nuclear equation if Iodine-131 undergoes β^- decay. Iodine has an atomic number of 53 and loses an electron to form an ion. How many protons and electrons are there? Write a nuclear equation if Potassium-37 undergoes β^- decay.

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R Sandford



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Table of Contents Zgce Inter Science Paper1 Revision

1. Understanding the eBook Zgce Inter Science Paper1 Revision
 - The Rise of Digital Reading Zgce Inter Science Paper1 Revision
 - Advantages of eBooks Over Traditional Books
2. Identifying Zgce Inter Science Paper1 Revision
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Zgce Inter Science Paper1 Revision
 - User-Friendly Interface
4. Exploring eBook Recommendations from Zgce Inter Science Paper1 Revision

- Personalized Recommendations
- Zgce Inter Science Paper1 Revision User Reviews and Ratings
- Zgce Inter Science Paper1 Revision and Bestseller Lists
- 5. Accessing Zgce Inter Science Paper1 Revision Free and Paid eBooks
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 - Zgce Inter Science Paper1 Revision eBook Subscription Services
 - Zgce Inter Science Paper1 Revision Budget-Friendly Options
- 6. Navigating Zgce Inter Science Paper1 Revision eBook Formats
 - ePub, PDF, MOBI, and More
 - Zgce Inter Science Paper1 Revision Compatibility with Devices
 - Zgce Inter Science Paper1 Revision Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Zgce Inter Science Paper1 Revision
 - Highlighting and Note-Taking Zgce Inter Science Paper1 Revision
 - Interactive Elements Zgce Inter Science Paper1 Revision
- 8. Staying Engaged with Zgce Inter Science Paper1 Revision
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Zgce Inter Science Paper1 Revision
- 9. Balancing eBooks and Physical Books Zgce Inter Science Paper1 Revision
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Zgce Inter Science Paper1 Revision
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Zgce Inter Science Paper1 Revision
 - Setting Reading Goals Zgce Inter Science Paper1 Revision
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Zgce Inter Science Paper1 Revision

- Fact-Checking eBook Content of Zgce Inter Science Paper1 Revision
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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