

Multiple Choice: (from Zumdahl *Chemistry*, 9th Edition, AP Edition)

Refer to the following diagram for questions 1 – 3

A 100.0 mL Graduated Cylinder is filled with 100.0 mL of water. A few drops of food colouring are added to the top of the water, without mixing. Initially, the top of the graduated cylinder appears to be very dark, while the bottom remains clear and colourless. After 5 min, the liquid in the graduated cylinder appears to be uniformly coloured.



1. Classify the contents in the graduated cylinder before any food colouring is added.

- (A) Element (B) Compound (C) Heterogeneous mixture
(D) Homogeneous mixture

2. Classify the contents in the graduated cylinder just after the food colouring is added.

- (A) Element (B) Compound (C) Heterogeneous Mixture (D) Homogeneous mixture

3. Classify the contents in the graduated cylinder after the 5 minutes has elapsed.

- (A) Element (B) Compound (C) Heterogeneous Mixture (D) Homogeneous mixture

4. A 10.00 g piece of metal is submerged in a graduated cylinder initially containing 20.00 mL of water. With the metal in the cylinder, the level of water is recorded as 21.40 mL. Determine the identity of the metal.

- (A) Al ($\rho = 2.70$ g/mL)
(B) Zn ($\rho = 7.13$ g/mL)
(C) Ag ($\rho = 10.49$ g/mL)
(D) Au ($\rho = 19.32$ g/mL)

5. $^{238}_{94}\text{Pu} \rightarrow \text{X} + \alpha$

In the decay represented above, what is the nucleon number of particle X?

- (A) 92 (B) 96 (C) 235 (D) 243

6. Which of the following statements about atoms is CORRECT?

- (A) Atoms are charged because they have different numbers of protons and neutrons
(B) All atoms of an element have the same number of nucleons and electrons.
(C) An atom is mostly filled with protons.
(D) The nucleus is negatively charged.

6. When each of the following is heated to 50°C and the temperature is held constant for 5 min, which one undergoes **only** a physical change.

- (A) egg (B) steak (C) ice cream (D) cake batter

7. A sample of Ti was ionized by electron impact in a mass spectrometer. Information from the mass spectrum about the isotopes of Ti in the sample is shown in Table 2.

Calculate the relative atomic mass of Ti in this sample. Report answers to one decimal place.

Table 1. Results from mass spectrum of Ti sample

| m/z | 46 | 47 | 48 | 49 |
|-------------------|-----|-----|------|-----|
| Percent Abundance | 9.1 | 7.8 | 74.6 | 8.5 |

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 - The Rise of Digital Reading Zumdahl Chemistry 9th Edition Multiple Choice
 - Advantages of eBooks Over Traditional Books
2. Identifying Zumdahl Chemistry 9th Edition Multiple Choice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Zumdahl Chemistry 9th Edition Multiple Choice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Zumdahl Chemistry 9th Edition Multiple Choice
 - Personalized Recommendations
 - Zumdahl Chemistry 9th Edition Multiple Choice User Reviews and Ratings
 - Zumdahl Chemistry 9th Edition Multiple Choice and Bestseller Lists
5. Accessing Zumdahl Chemistry 9th Edition Multiple Choice Free and Paid eBooks
 - Zumdahl Chemistry 9th Edition Multiple Choice Public Domain eBooks
 - Zumdahl Chemistry 9th Edition Multiple Choice eBook Subscription Services
 - Zumdahl Chemistry 9th Edition Multiple Choice Budget-Friendly Options
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- Zumdahl Chemistry 9th Edition Multiple Choice Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Zumdahl Chemistry 9th Edition Multiple Choice
 - Highlighting and Note-Taking Zumdahl Chemistry 9th Edition Multiple Choice
 - Interactive Elements Zumdahl Chemistry 9th Edition Multiple Choice
- 8. Staying Engaged with Zumdahl Chemistry 9th Edition Multiple Choice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Zumdahl Chemistry 9th Edition Multiple Choice
- 9. Balancing eBooks and Physical Books Zumdahl Chemistry 9th Edition Multiple Choice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Zumdahl Chemistry 9th Edition Multiple Choice
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Zumdahl Chemistry 9th Edition Multiple Choice
 - Setting Reading Goals Zumdahl Chemistry 9th Edition Multiple Choice
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Zumdahl Chemistry 9th Edition Multiple Choice
 - Fact-Checking eBook Content of Zumdahl Chemistry 9th Edition Multiple Choice
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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