

8.2 The Nature of Covalent Bonding

Essential Understanding

Covalent bonds form when atoms share electrons.

Lesson Summary

The Octet Rule in Covalent Bonding

Covalent compounds are most stable when each atom has eight electrons.

- ▶ Single, double, and triple covalent bonds depend on the number of pairs of electrons shared between two atoms.
- ▶ Atoms form double or triple covalent bonds if they can attain a noble gas structure by doing so.

Type of Covalent Bond	Attributes
Single	One shared electron pair with one electron from each atom
Double	Two shared electron pairs with two electrons from each atom
Triple	Three shared electron pairs with three electrons from each atom

Coordinate Covalent Bonds In a coordinate covalent bond, one atom contributes both electrons in the bonding pair.

- ▶ One atom may contribute a pair of unshared electrons to a bond to give both atoms an inert gas configuration.
- ▶ Coordinate covalent bonds can also occur in polyatomic ions, such as NH_4^+ .

Exceptions to the Octet Rule Some molecules have fewer, or more, than a complete octet of valence electrons.

- ▶ Molecules that have an odd number of total valence electrons cannot satisfy the octet rule.
- ▶ Some molecules that have an even number of valence electrons may also fail to follow the octet rule.

The Nature Of Covalent Bonding 16 Worksheet

Anil N. Netravali, K. L. Mittal



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Kresan, Reed Mencke, Frank Press, Raymond Siever, 2003-09-25 This reconceptualization of the text *Understanding Earth* reflects the fundamental changes in the field of physical geology over the past several years. **Functional Tactile**

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Thomas,2023-04-29 This book on biopolymers offers a comprehensive source for biomaterial professionals It covers all elementary topics related to the properties of biopolymers the production and processing of biopolymers applications of biopolymers examples of biopolymers and the future of biopolymers Edited by experts in the field the book highlights international professionals longstanding experiences and addresses the requirements of practitioners and newcomers in this

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computational modelling and numerical simulation tools and confirmed by modern experimental techniques It begins by summarizing basic theoretical methods then giving both a theoretical and experimental treatment of how alkali metal clusters develop into nanostructures as influenced by the cluster's magic number of atoms The book continues with a discussion of atomic clusters and nanostructures focusing primarily on boron and carbon exploring in detail the one two and three dimensional structures of boron and carbon and describing their myriad potential applications in nanotechnology from nanocoating and nanosensing to nanobatteries with high borophene capacity The broad discussion of computational modelling as well as the specific applications to boron and carbon make this book an essential reference resource for materials scientists in this field of research

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