

THE OLD TESTAMENT				
5	12	5	5	12
<b>Pentateuch</b>	<b>History</b>	<b>Wisdom &amp; Poetry</b>	<b>Major Prophets</b>	<b>Minor Prophets</b>
Genesis Exodus Leviticus Numbers Deuteronomy	Joshua Judges Ruth 1 & 2 Samuel 1 & 2 Kings 1 & 2 Chronicles Ezra Nehemiah Esther	Job Psalms Proverbs Ecclesiastes Song of Songs	Isaiah Jeremiah Lamentations Ezekiel Daniel	Hosea Joel Amos Obadiah Jonah Micah Nahum Habakkuk Zephaniah Haggai Zechariah Malachi

12

5

12

## History

## Major Prophets

## Minor Prophets

Joshua  
Judges  
Ruth  
1 & 2 Samuel  
1 & 2 Kings  
1 & 2  
Chronicles  
Ezra  
Nehemiah  
Esther

Isaiah  
Jeremiah  
Lamentations  
Ezekiel  
Daniel

Hosea  
Joel  
Amos  
Obadiah  
Jonah  
Micah  
Nahum  
Habakkuk  
Zephaniah  
Haggai  
Zechariah  
Malachi

# The Old Testament Prophets An Outline

**George Mark Elliott**



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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3\text{ mole H}_2\text{S} \times \frac{2\text{ mole H}_2\text{O}}{2\text{ mole H}_2\text{S}} \times 18\text{g H}_2\text{O} = 149.4\text{g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}_2$  what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3 \text{ g}}{32 \text{ g/mol}} = 0.666 \text{ mol}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 160.0 \text{ g}$  c determine the percent yield  $\frac{126.0 \text{ g}}{160.0 \text{ g}} \times 100 = 78.75\%$

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web 7 Na  $2\text{C}_2\text{O}_4$  is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  xs 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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