# Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON METALS share electrons, attempting to attain a stable octet (8 outer electrons)in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

H<sub>2 brothrogen</sub> is distornic

H + H → H → H-H Single Bond

2.) F2 thurston in discounts

3.) O2 oxygen is diatomic

O + O

 $\bullet$   $\bigcirc =\bigcirc$ 

Double Bond

4.) No nitrogen is diatomic. Is this a triple bond?

N + N

5.) BF<sub>3 you need 3 fluorine atoms bere-</sub>

 $\mathbf{R} + \mathbf{F}$ 

6.) Ammonia NH3 hist: how many hydrogen atoms are needed?

N + H

7.) Carbon dioxide CO<sub>2</sub>

0 + c + 0

Methane CH<sub>4</sub> careful here 4 hydrogen atoms needed

C + H

9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!

H + O + H

10.) SO<sub>2</sub> hint: one pair of electrons from sulfur must be slit up for this one to work.

O + S + O

# **The Octet Rule In Covalent Bonding Worksheet Answers**

N. D. Epiotis

#### The Octet Rule In Covalent Bonding Worksheet Answers:

Holt Chemistry Ralph Thomas Myers, 2004 Holt Chemistry Holt Rinehart & Winston, 2003-01-24 Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1963 Molecules and the Chemical Bond Henry A. Bent, 2011 MOLECULES AND THE CHEMICAL BOND Chemistry Simplified This highly original book by a famous chemistry teacher about general chemistry in a new key may change how teachers teach Atomic Theory The Mole Concept and Avogadro's Constant The Gas Laws Solving Problems in Chemical Stoichiometry The Saturation and Directional Character of Chemical Affinity The Pauli Exclusion Principle Linnett's Double Spin Set Theory Pauling's Rules of Crystal Chemistry The Octet Rule Lewis Structures for O2 NO CO SO2 and SO3 Construction of Bond Diagrams VSEPR Theory Dative Bonding Multicenter Bonding Bonding in Metals pH Calculations The Periodic Table The Energy Function and the First Law of Thermodynamics The Entropy Function and the Second Law of Thermodynamics How an Inductive Science Advances

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