

Mole Conversions Practice Answer Key

Molar Mass – Calculate the Molar Mass for the following.

- What is the molar mass of Magnesium metal?

$$24.305 \text{ g/mol}$$

- Calculate the molar mass of NaCl

$$\begin{aligned} \text{Na} &: 22.990 \times 1 = 22.990 \\ \text{Cl} &: 35.453 \times 1 = 35.453 \end{aligned} \Rightarrow 58.443 \text{ g/mol}$$

- What is the molar mass of NaOH?

$$\begin{aligned} \text{Na} &: 1 \times 22.990 = 22.990 \\ \text{O} &: 1 \times 15.999 = 15.999 \\ \text{H} &: 1 \times 1.008 = 1.008 \end{aligned} \Rightarrow 39.997 \text{ g/mol}$$

Conversions – Calculate the following the conversions. Use your Mole Map to help you!

- Determine the number of atoms in 0.50 moles He.

$$0.5 \text{ mol He} \times \frac{6.02 \times 10^{23} \text{ atoms He}}{1 \text{ mol He}} = 3.01 \times 10^{23} \text{ atoms He}$$

- Determine the number of moles in 80.00 g O₂

$$0.2 \times 15.999 = 31.998 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol} = 31.998 \text{ g}$$

- Determine the volume of 0.24 mol of NO₂.

$$0.24 \text{ mol NO}_2 \times \frac{22.4 \text{ L NO}_2}{1 \text{ mol NO}_2} = 5.376 \text{ L NO}_2$$

$$80.00 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 25.00 \text{ g O}_2$$

- Determine the number of moles in 7.5×10^{19} formula units of KNO₃

$$7.5 \times 10^{19} \text{ f.u. KNO}_3 \times \frac{1 \text{ mol KNO}_3}{6.02 \times 10^{23} \text{ f.u. KNO}_3} = 1.2458 \times 10^{-4} \text{ f.u. KNO}_3$$

- Determine the number of moles in 202.22g KNO₃

$$\begin{aligned} \text{K} &: 1 \times 39.098 = 39.098 \\ \text{N} &: 1 \times 14.007 = 14.007 \\ \text{O} &: 3 \times 15.999 = 47.997 \end{aligned} \Rightarrow 101.102 \text{ g/mol} \Rightarrow 1 \text{ mol} = 101.102 \text{ g}$$

- Determine the number of moles in 32.5L of N₂ at STP

$$32.5 \text{ L N}_2 \times \frac{1 \text{ mol N}_2}{22.4 \text{ L N}_2} = 1.451 \text{ mol N}_2$$

$$202.22 \text{ g KNO}_3 \times \frac{1 \text{ mol KNO}_3}{101.102 \text{ g KNO}_3} = 2.00 \text{ mol KNO}_3$$

- Determine the number of moles in 3.4×10^{26} formula units of NaCl.

$$3.4 \times 10^{26} \text{ f.u. NaCl} \times \frac{1 \text{ mol NaCl}}{6.02 \times 10^{23} \text{ f.u. NaCl}} = 569.784 \text{ f.u. NaCl}$$

- Determine the moles in 36.5L of carbon monoxide

$$36.5 \text{ L CO} \times \frac{1 \text{ mol CO}}{22.4 \text{ L CO}} = 1.629 \text{ mol CO}$$

The Mole Chapter Answer Key

Rachel Sandford



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asterix and obelix s birthday the golden book french asterix et obelix l anniversaire le livre d or is the thirty fourth episode of the asterix comic book series designed and written by albert uderzo the usage of previous asterix stories led to co creator rené goscinnny also being credited the album which celebrates the 50th

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