

# The Mole Study Guide

**Mass Composition:** What's your element's % of mass after the reaction? Molar mass to obtain your value!

10. There are 3.75 moles of  $\text{CO}_2$ .

11. There are 3.75 moles of  $\text{CO}_2$ .

12. There are 3.75 moles of  $\text{CO}_2$ .

13. There are 3.75 moles of  $\text{CO}_2$ .

14. There are 3.75 moles of  $\text{CO}_2$ .

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24. There are 3.75 moles of  $\text{CO}_2$ .

25. There are 3.75 moles of  $\text{CO}_2$ .

26. There are 3.75 moles of  $\text{CO}_2$ .

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## The Mole Study Guide

Fill in the Blank boxes for the image below.



Write the Empirical Formula for each compound. Use subscripts to show the relative ratio of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_5\text{H}_{10}\text{O}_5$	

Calculate the Molar Mass for the following compounds.

10.  $\text{H}_2\text{O}$

11.  $\text{H}_2\text{O}_2$

12.  $\text{H}_2\text{O}$

13.  $\text{H}_2\text{O}_2$

Calculate the Percent Composition for each element in the following compound.

10.  $\text{H}_2\text{O}$

11.  $\text{H}_2\text{O}_2$

12.  $\text{H}_2\text{O}$

13.  $\text{H}_2\text{O}_2$

14.  $\text{H}_2\text{O}$

15.  $\text{H}_2\text{O}_2$

16.  $\text{H}_2\text{O}$

17.  $\text{H}_2\text{O}_2$

18.  $\text{H}_2\text{O}$

19.  $\text{H}_2\text{O}_2$

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 40.00% Hg, 21.87% Cl, 3.12% N, and 34.91% O. The molar mass for Astatine is 210.0 g/mol.

a. If you have 100 g of Astatine, how many grams of Hg, Cl, and O do you have?

b. Determine the Empirical Formula for Astatine.

c. Determine the Molecular Formula for Astatine.

# The Mole Study Guide

**Intelligent Education**



## **The Mole Study Guide:**

*The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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