

8.2 The Nature of Covalent Bonding

Connecting to Your World

You know that without oxygen to breathe, you could not live. But did you know that oxygen plays another important role in your life? High in the atmosphere, a different form of oxygen, called ozone, forms a layer that filters out harmful radiation from the sun. The colors in this map indicate the concentrations of ozone in various parts of Earth's atmosphere. In this section, you will learn how oxygen atoms can join in pairs to form the oxygen you breathe and can also join in groups of three oxygen atoms to form ozone.



The Octet Rule in Covalent Bonding

Recall that when ionic compounds form, electrons tend to be transferred so that each ion acquires a noble gas configuration. A similar rule applies for covalent bonds. **In covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of noble gases.** For example, each hydrogen atom has one electron. But a pair of hydrogen atoms share these two electrons when they form a covalent bond in a hydrogen molecule. Each hydrogen atom thus attains the electron configuration of helium, a noble gas with two electrons. Combinations of atoms of the nonmetals and metalloids in Groups 4A, 5A, 6A, and 7A of the periodic table are likely to form covalent bonds. In this case the atoms usually acquire a total of eight electrons, or an octet, by sharing electrons, so that the octet rule applies.

Single Covalent Bonds

The hydrogen atoms in a hydrogen molecule are held together mainly by the attraction of the shared electrons to the positive nuclei. Two atoms held together by sharing a pair of electrons are joined by a **single covalent bond**. Hydrogen gas consists of diatomic molecules whose atoms share only one pair of electrons, forming a single covalent bond.



Guide for Reading

Key Concepts

- What is the result of electron sharing in covalent bonds?
- How do electron-dot structures represent shared electrons?
- How do atoms form double or triple covalent bonds?
- How are coordinate covalent bonds different from other covalent bonds?
- How is the strength of a covalent bond related to its bond dissociation energy?
- How are oxygen atoms bonded in ozone?
- What are some exceptions to the octet rule?

Vocabulary

single covalent bond
structural formula
unshared pair
double covalent bond
triple covalent bond
coordinate covalent bond
polyatomic ion
bond dissociation energy
resonance structure

Reading Strategy

Identifying Main Ideas/Details

List the main idea in the paragraph following the heading **The Octet Rule in Covalent Bonding**. As you read, list examples of how this rule is followed when a single covalent bond, a double covalent bond, and a triple covalent bond form.

The Nature Of Covalent Bonding 8 2 Answers

Rachel Sandford



The Nature Of Covalent Bonding 8 2 Answers:

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different aspects of Inorganic Chemistry in terms of 10 Chapters with in depth and up to date coverage Starting with the VSEPR theory in the first chapter the book symmetrically presents delocalized p bonding in polyatomic molecules structure bonding and topology of borane and related compounds synthesis and reactivity of metal clusters and their bonding some aspects of stability constants of metal complexes magnetochemistry mechanism of inorganic reactions molecular orbital MO approach of bonding in transition metals bonding in organometallic sandwich compounds based on MO approach Safe and economical inorganic experiments at UG and PG Levels are also presented in the last chapter At the end five relevant topics are included as appendices for updating students and faculty members

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Molecular Biology of the Cell 6E - The Problems Book John Wilson, Tim Hunt, 2014-11-21 The Problems Book helps students appreciate the ways in which experiments and simple calculations can lead to an understanding of how cells work by introducing the experimental foundation of cell and molecular biology Each chapter reviews key terms tests for understanding basic concepts and poses research based problems The Problems Book has been

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