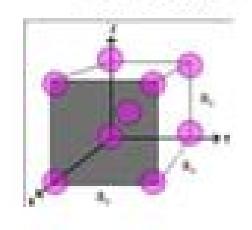
Solution for plane (100)

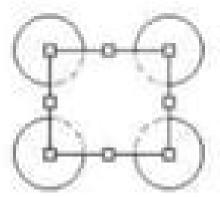
First, find atomic radius for Nickel from Appendix 2, page 797 of Textbook (6th Ed. Shackleford) to calculate the lattice parameter:



$$r_{ij} = 0.152nm$$

$$a_0(BCC) = \frac{4r}{\sqrt{3}} = \frac{4(0.152nm)}{\sqrt{3}} = 0.3510nm \left(3.510\text{ Å}\right)$$

For (100):

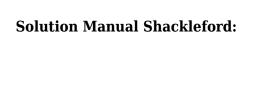


$$planar_density = \frac{latom}{(3.510 \times 10^{-6} cm)^2} = 8.115 \times 10^{14} atoms / cm^2$$

$$packing_fraction = \frac{\pi \left(\sqrt{3}a_0/4\right)^2}{a_0^2} = 0.5890$$

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