



## Step 2

2 of 5

At  $t = 2s$

$$\begin{aligned} v_{1\Omega} &= 1.8 \sin(4) \\ &= 0.1256V \end{aligned} \quad (1)$$

$$P_{1\Omega} = \frac{(v_{1\Omega})^2}{R} \quad (2)$$

$$P_{1\Omega} = \frac{(0.1256)^2}{1} \quad (3)$$

$$P_{1\Omega} = 0.01577w \quad (4)$$

$$\boxed{P_{1\Omega} = 15.77mW} \quad (5)$$

## Step 3

3 of 5

c)  $v_s = 9 \sin(2t + 13^\circ) \text{ V}$

$$v_{1\Omega} = \frac{[9 \sin(2t + 13^\circ)] \cdot 1}{5} \quad (1)$$

$$= 1.8 \sin(2t + 13^\circ) \quad (2)$$

At  $t = 0$

$$\begin{aligned} v_{1\Omega} &= 1.8 \sin(13^\circ) \\ &= 0.4049 \text{ V} \end{aligned} \quad (3)$$

$$P = \frac{(v_{1\Omega})^2}{1} \quad (4)$$

$$= \frac{0.4049^2}{1} \quad (5)$$

$$= 0.164W \quad (6)$$

$$\boxed{P = 0.164W} \quad (7)$$

At  $t = 1s$

$$\begin{aligned} v_{1\Omega} &= 1.8 \sin(2 + 13) \\ &= 0.4659 \end{aligned} \quad (8)$$

$$P = \frac{(v_{1\Omega})^2}{1} \quad (9)$$

$$= \frac{0.4659^2}{1} \quad (10)$$

$$= 0.217W \quad (11)$$

$$\boxed{P = 0.217W} \quad (12)$$

At  $t = 2s$

$$\begin{aligned} v_{1\Omega} &= 1.8 \sin(4 + 13) \\ &= 0.5263V \end{aligned} \quad (13)$$

$$P = \frac{(v_{1\Omega})^2}{1} \quad (14)$$

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**B. S. Kesavan**



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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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eg 1 hcl aq + naoh aq → nacl aq + h<sub>2</sub>o l for which Δh = 57 kJ mol<sup>-1</sup> eg 2 h<sub>2</sub>so<sub>4</sub> aq + 2 kno<sub>3</sub> aq + h<sub>2</sub>o l for which Δh = 57 kJ mol<sup>-1</sup>

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