

Solutions, acids, and bases

1. VOCABULARY REVIEW OF SOLUTIONS

Solute - what is **dissolved** in the solvent to form a solution.

Solvent - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

Miscible - two mix in any proportions (none to be mixed) - liquids will pour each over

Immiscible - don't mix in any proportions (none to be mixed)

Examples of miscible-immiscible - ethanol-water, oil-water, oil-water, oil-water

Solute - **amount** of **being dissolved** in liquid

Insoluble - a substance (tends to be a solid) that **will not dissolve** in a solvent (an example is sand and water)

Suspension - looks like a **solution** when **stirred**, the **solutes** a bit of water (suspensions) also it is a **hetero** - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

Unsaturated - **solids** with **solubility chart** when the temperature would be enough to make a substance dissolve but at a greater amount of **solute** - **solubility less** **solute** **dissolves** than can (ex. salt @ 20°C temp)

Saturated - **solids** with **solubility chart** when the temperature - **solubility** relationship is at the line in **over** words a **solution** with **solute** has **reached** at the **maximum possible amount**.

Supersaturated - contains **more** than the **max.** amount of **solute** that can be dissolved at the temperature on the **solubility chart** this would mean the grams of solute would go above the line.

2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **supersaturated**

At the line - **saturated**

Under the line - **unsaturated**

3. METHODS OF SEPARATION

Filtration - separate particles by **size**

Distillation - separate by **boiling point** - **more** **volatile** **liquids** to be separated.

Extraction - allows one to remove a **liquid** from a **solid** by **transferring** the **liquid** to a **gaseous phase**.

Centrifugation - separating molecules by their densities by spinning them apart.

Decanting - process of separating insoluble or suspension solutions by **pouring** away the liquid from the solid.

Chromatography - separation by the **rates** at which components of different solutions move through a medium (ex. gel electrophoresis)

Solutions Acids And Bases Study Guide

**John Krenos, Joseph Potenza, Carl
Hoeger**



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