

## Chemical Equilibrium

**Ex 1.** For the equilibrium:  $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$

the equilibrium constant  $K_c$  is 50.5 at 448°C.

If  $[\text{H}_2]_0 = [\text{I}_2]_0 = 1 \text{ mol.L}^{-1}$ , what will the concentration of  $\text{H}_2$ ,  $\text{I}_2$  and  $\text{HI}$  be at equilibrium.

**Ex 2.** Consider the following reversible reaction:



One mole of  $\text{PCl}_5$  is placed in a flask of capacity 4 liter to be dissociated at constant temperature T. the number of moles of  $\text{Cl}_2$  formed at equilibrium is 0.25mole.

Calculate the equilibrium constant of the above system at temperature T.

**Ex 3.** At 100°C, the equilibrium constant for the reaction system,



is  $K_c = 2.19 \times 10^{-10}$

1. Determine whether the following mixture of components of the above system is at equilibrium:

$$[\text{COCl}_2] = 3.5 \times 10^{-3} \text{ mol.L}^{-1}$$

$$[\text{CO}] = 1.2 \times 10^{-5} \text{ mol.L}^{-1}$$

$$[\text{Cl}_2] = 3.3 \times 10^{-6} \text{ mol.L}^{-1}$$

2. If it is not at equilibrium, in what direction the system proceeds to reach equilibrium?

**Ex 4.** In a closed container of volume V and at constant temperature  $T = 300\text{K}$ , we introduce 2 mol of NO and 1 mol of  $\text{O}_2$  gases. the following equilibrium is established:



Calculate the degree of conversion ( $\alpha$ ) of NO if the amount of  $\text{NO}_2$  at equilibrium is three time that of  $\text{O}_2$ .

**Ex 5.** Ethanamine  $\text{C}_2\text{H}_5\text{NH}_2$  gas is put in an empty flask of volume  $V = 2\text{L}$ , maintained at constant temperature  $T = 500\text{K}$ .

The following equilibrium is established:



The initial amount of ethanamine is 0.2 mol and the total amount of the mixture at equilibrium is 0.3 mol.

1. Calculate the degree of dissociation  $\alpha$  of ethanamine.
2. Calculate the equilibrium constant  $K_c$  at 500K.

# Solutions Manual Chemical Equilibrium Chemistry

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## **Solutions Manual Chemical Equilibrium Chemistry:**

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*Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition* Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools **Student's Solutions Manual to Accompany Atkins' Physical Chemistry** C. A. Trapp, M. P. Cady, Carmen

Giunta, 2010 This solutions manual provides the authors detailed solutions to exercises and problems in physical chemistry It comprises solutions to exercises at the end of each chapter and solutions to numerical theoretical and additional problems

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