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Solutions, Acids, and Bases Test Review

Solutions

- 1- In soda pop, $\text{CO}_2(\text{g})$ is a **solute** and water is the **solvent**.
- 2- What term is used by chemists to quantitatively describe a solution in which a relatively small amount of solute is dissolved? **Dilute/unsaturated**
- 3- When a solvent has dissolved all the solute it can at a particular temperature, it is said to be _____. **Saturated**
- 4- What will increase the rate of dissolution of a solute in a solvent? **Stirring, increasing surface area, increasing temperature**
- 5- What is the mass of H_2SO_4 in 1.00×10^2 mL of 0.200 M H_2SO_4 solution? **1.96g**
- 6- What number of moles of solute are present in 25.0 mL of 2.00 M HCl? **0.05 Moles**
- 7- How many grams of K_2SO_4 (molar mass = 174 g/mol) would be needed to prepare 4.00 L of a 0.0510 M solution? **35.5g**
- 8- A solution is prepared by dissolving 7.31 g of Na_2SO_4 in enough water to make 225 mL of solution. Calculate the solution molarity. **0.229M**
- 9- A 54.8 g sample of SrCl_2 is dissolved in 112.5 mL of solution. Calculate the molarity of this solution. **3.07M**

Gas Laws

- 1- A balloon has a volume of 1.20 L at 24.0°C . The balloon is heated to 48.0°C . Calculate the new volume of the balloon. **1.30L**
- 2- A sample of helium gas occupies 2.65 L at 1.20 atm. What pressure would this sample of gas exert in a 1.50 L container at the same temperature? **2.12atm**
- 3- A gas has a volume of 5.0 L at a certain pressure. How must the pressure be changed to double the volume of the gas at constant temperature? **Pressure is 1/2**
- 4- A 6.5 L sample of nitrogen at 25°C and 1.5 atm is allowed to expand to 13.0 L. The temperature remains constant. What is the final pressure? **0.75atm**
- 5- You transfer a sample of a gas at 17°C from a volume of 5.67 L and 1.10 atm to a container at 37°C that has a pressure of 1.10 atm. What is the new volume of the gas? **6.06L**
- 6- A sample of helium gas occupies 12.4 L at 23°C and 0.956 atm. What volume will it occupy at 40°C and 0.956 atm? **13.1L**
- 7- A sample of oxygen gas (O_2) has a volume of 7.51 L at a temperature of 19°C and a pressure of 1.38 atm. Calculate the moles of O_2 molecules present in this gas sample. **0.433 moles**
- 8- A weather balloon contains helium gas at a volume of 855 L. Initially, the balloon is at sea level, where the temperature is 25°C and the barometric pressure is 0.947 atm. The balloon then rises to an altitude of 6000 ft, where the pressure is 0.796 atm and the temperature is 15°C . What is the change in the volume of the balloon as it ascends from sea level to 6000 ft? **128L**
- 9- A 6.35 L sample of carbon monoxide is collected at 55°C and 0.892 atm. What volume will the gas occupy at 1.05 atm and $20.^\circ\text{C}$? **4.82L**

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